

Computational Study of Medium-Sized Cumulenones, H_2C_nO ($n = 3-7$)

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Computations are reported for the very floppy cumulenones H_2C_nO ($n = 3-7$). The structures, harmonic frequencies and dipole moments are computed using the DFT and MP2 methods. H_2C_6O and H_2C_7O are treated for the first time. The carbon skeletons in ground states of these molecules with smaller number of carbon atoms are found to be bent, but very floppy. Those of the cumulenones with higher number of carbon atoms are predicted to be nearly linear. All the molecules are predicted to be of C_s symmetry by the present results. Good agreement with the available experimental observations is obtained.

Introduction

Carbon clusters with polycumulated double bonds of general structure of XC_nY ($X, Y = H_2, O, S$) received considerable attention recently due to the unusual spectroscopic properties and high reactivity. Some of these molecules are known to be very quasilinear and very floppy.¹ For example, the frequency of the lowest bending mode of the linear molecule OC_3O is only 18 cm^{-1} .² Determination of the structures and spectroscopic properties of these molecules are considered challenging, due to the large size and the floppiness of the molecules, which are also of astrophysical interests.

Considering the spectroscopic and astrophysical importance of these molecules, we report in the present work the theoretical calculations by employing the density functional theory (DFT),^{3,4} which proved⁵⁻⁹ to give surprisingly accurate spectroscopic properties of medium-sized carbon clusters with much less computational efforts than other correlated *ab initio* methods. The effectiveness of the DFT stems from the fact that computational efforts of the method increase as n^3 (n is the number of electrons) with comparable accuracy to the more conventional post-HF techniques, while those of the latter methods are proportional as $n^4 \sim n^7$.¹⁰ In this paper we report the spectroscopic properties of H_2C_nO ($n=3-7$). We employ the BLYP, B3LYP and MP2 methods employing basis sets up to 6-311G**.

Computational Methods

In this study all the calculations were carried out using the GAUSSIAN 94 set of programs.¹¹ We find that the density functional theory with the exchange functional of Becke¹² and correlation functional by Lee *et al.*¹³ (BLYP) gives the best overall results. Bond lengths, angles, rotational constants and dipole moments are reported along with the zero-point energies and harmonic frequencies.

Results

The energies, rotational constants, dipole moments, and the zero point energies of the ground states of linear H_2C_nO ($n = 3-7$) are presented in Table 1. In their lowest energy structures, the electronic states are all singlet. The computed rotational constants and dipole moments of H_2C_3O and H_2C_4O agree well with experimental observations¹⁴⁻¹⁷ except one rotational constant (A) about the main axis. For a quasilinear molecule, it is difficult to predict A value which approaches to infinity. Calculated bond lengths and harmonic frequencies are reported in Table 3-Table 7.

Since the focus of the present computations is whether the molecules are of C_{2v} symmetry or not, we present the angles of the $C_\gamma-C_\beta-C_\alpha$ bonds adjacent to the oxygen atom (we will call it the "kink" angle here) in Table 2. All the computa-

Table 1. Total energies (hartree), zero-point energies E_0 (kcal/mol), rotational constants B_e (GHz), and dipole moments μ (debye) of the ground states of H_2C_nO ($n=3-7$) by BLYP/6-311G**

n	Energy	E_0	B_e		μ	
			Cal.	Exp.	Cal.	Exp.
3	-190.67176	22.6843	164.2, 4.290, 4.180	153.5, 4.387, 4.258 ^a	2.3470	2.297 ^b
4	-228.76396	25.2646	280.6, 2.125, 2.110	220, 2.16, 2.15 ^c	1.9734	1.967 ^d
5	-266.83571	28.4957	271.4, 1.234, 1.228	N/A	3.1054	N/A
6	-304.92459	31.6218	287.8, 0.785, 0.783	N/A	2.5372	N/A
7	-342.99985	34.9614	289.6, 0.529, 0.528	N/A	3.6508	N/A

^aexperimental values; Ref. 14., ^bexperimental values; Ref. 15. ^cexperimental values; Ref. 16., For a quasilinear molecule, B(along z-axis) has infinite value. ^dexperimental values; Ref. 17.

Table 2. "Kink" angles (in degrees) of $C_\gamma-C_\beta-C_\alpha$ for H_2C_nO ($n = 3-7$)

n	MP2/ 6-311G**	B3LYP/ 6-311G**	BLYP/ 6-311G**	Prediction ^a
3	137.7	153.1	148.1	145 ± 3
4	156.2	176.5	176.0	172 ± 8
5	150.1	173.9	174.4	172 ± 8
6	166.5	178.6	179.9	N/A
7	159.3	176.2	179.9	N/A

^aRef. 18.

tional methods employed yield some degree of kinkiness to the molecules. Comparison is also made with the very recent theoretical predictions of East,¹⁸ which are based on computations employing MP2, MP4, QCISD and QCISD(T) methods with 6-31G(d), cc-pVDZ, and augmented cc-pVDZ basis sets. The kinky angles obtained by MP2 method seem to be too small compared with other computed results. The C-C-C bond of H_2C_3O seems to be highly bent with the kinky angle of 140-150°. Since the canonical structure of H_2C_3O , $H_2C_\gamma=C_\beta=C_\alpha=O$, would give C_{2v} structure, there seems to be strong contribution of other resonance structure, notably $H_2C_\gamma=C_\beta-C_\alpha\equiv O$, in which the C_β atom has a lone pair, as proposed by East.¹⁸ For H_2C_nO with n larger than 4, the kinky angle of $C_\gamma-C_\beta-C_\alpha$ bond is larger than 120°, indicating that the molecules are nearly of C_{2v} symmetry. The kink angles of H_2C_6O and H_2C_7O , which are treated for the first time, are predicted to be nearly 180° by the present computations. For H_2C_nO ($n \geq 4$), the C_δ atom may also possess a lone pair, and the overall importance of the resonance

structures with the $C_\delta=C_\gamma=C_\beta-C_\alpha\equiv O$ moiety, with lone pair at C_β , may decrease, giving structures closer to the C_{2v} structures. Accordingly, the $C_\gamma=C_\beta-C_\alpha$ bending potential will flatten, and the harmonic frequency of this mode will also decrease, as we will show below.

Table 3-Table 7 present the computed bond angles and bond lengths of the optimized singlet ground states of H_2C_nO ($n = 3-7$), along with the available experimental data. The C-C bond adjacent to the methylene group is much longer than other C-C bonds in all the molecules studied, as also was predicted by East. Agreement of our computed results with the experimental data is very good. It can be seen that the methylene angles (H-C-H) are less than 120° for all the cumulenones. Lengths of the C=O bond and the C=C bond adjacent to the methylene carbon exhibit oscillatory pattern with respect to the number of carbon atoms n . This pattern was also observed for some bond lengths and stretching frequencies in other carbon clusters bonded with heteroatoms.⁵⁻⁹

Table 3-Table 7 also give the harmonic frequencies of the cumulenones treated in this report. It should be noted that no scaling factors were employed here for the harmonic frequencies computed by BLYP and B3LYP to compare with experimental observations. The MP2 method overestimates the harmonic frequencies as usual, and appropriate scaling factors would have to be employed, as was done by East. Overall, the BLYP/6-311* or BLYP/6-311** methods give the results that are in the best agreement with experimental observations. For H_2C_3O , harmonic frequencies computed by the BLYP method excellently agree with experimental

Table 3. Structures (bond lengths in Å and bond angles in degrees) and harmonic frequencies (cm^{-1}) of H_2C_3O

Method	BLYP	BLYP	BLYP	B3LYP	B3LYP	B3LYP	MP2	MP2 ^a	Exp. ^b
Basis set	6-31G*	6-311G*	6-311G**	6-31G*	6-311G*	6-311G**	6-311G**	6-31G*	
R_{CH}	1.0983	1.0961	1.0953	1.0909	1.0893	1.0887	1.0885	1.0885	1.094
$R_{\beta\gamma}$	1.3400	1.3338	1.3335	1.3267	1.3215	1.3210	1.3401	1.3401	1.3219
$R_{\alpha\beta}$	1.3140	1.3034	1.3026	1.2986	1.2893	1.2886	1.3197	1.3200	1.3198
R_{CO}	1.1911	1.1820	1.1824	1.1781	1.1691	1.1695	1.1721	1.1720	1.1631
A_{HCH}	115.112	115.185	115.462	115.301	115.401	115.668	116.064	116.070	117.18
A_{HCC}	121.186	121.462	121.389	121.466	121.686	121.588	120.604	120.590	120.95
A_{HCC}	123.701	123.352	121.389	123.233	122.913	122.745	123.332	123.350	122.47
A_β	144.896	147.468	148.121	150.058	152.284	153.081	137.700	137.620	144.5
A_α	167.235	169.390	169.594	169.556	171.434	171.610	167.932	167.840	169.4
$\nu_9(A')$	172(15.5) [†]	163(17.8)	161(17.9)	139(18.9)	134(20.9)	131(20.8)	187(3.7)	205(17.6)	
$\nu_{12}(A'')$	264(2.5)	262(1.4)	263(1.2)	265(2.0)	261(0.8)	261(0.7)	274(2.1)	279(4.6)	
$\nu_8(A')$	463(12.0)	480(10.2)	478(10.2)	478(14.3)	503(15.0)	500(15.0)	504(11.6)	515(10.7)	479
$\nu_{11}(A'')$	677(3.6)	681(12.0)	681(12.7)	689(15.4)	696(19.4)	696(20.3)	686(8.9)	710(8.1)	674
$\nu_7(A')$	927(5.5)	919(4.1)	917(4.2)	955(3.8)	946(4.3)	945(4.3)	898(1.8)	961(1.6)	910
$\nu_{10}(A'')$	974(1.3)	974(25.9)	980(23.2)	1018(21.9)	1018(25.5)	1023(23.1)	962(25.2)	1022(27.4)	988
$\nu_6(A')$	1054(12.6)	1055(12.0)	1045(11.8)	1079(12.8)	1080(11.9)	1070(11.5)	1036(23.7)	1109(22.3)	1047
$\nu_5(A')$	1473(2.4)	1461(1.7)	1448(1.7)	1510(2.5)	1497(1.6)	1484(1.5)	1427(1.2)	1530(1.4)	1457
$\nu_4(A')$	1703(7.1)	1699(8.9)	1698(9.4)	1785(9.6)	1778(12.5)	1777(13.6)	1629(6.6)	1735(6.6)	1685
$\nu_3(A')$	2146(642.5)	2156(728.1)	2158(723.9)	2254(847.4)	2260(939.5)	2263(932.8)	2100(677.6)	2207(617.3)	2124
$\nu_2(A')$	3011(33.5)	2986(43.9)	2985(37.6)	3100(29.4)	3072(37.9)	3069(31.6)	2971(26.4)	3163(22.8)	2968.3
$\nu_1(A')$	3080(19.4)	3055(20.8)	3054(16.2)	3171(15.4)	3144(16.4)	3141(12.2)	3058(8.8)	3251(10.3)	3026.3

[†]IR intensities are in parentheses (KM/Mole). ^aRef. 18. ^bStructure data; Ref. 14., frequency data; Ref. 15.

Table 4. Structures (bond lengths in Å and bond angles in degrees) and harmonic frequencies (cm⁻¹) of H₂C₄O

Method	BLYP	BLYP	BLYP	B3LYP	B3LYP	B3LYP	MP2	MP2 ^a	Exp ^b
Basis set	6-31G*	6-311G*	6-311G**	6-31G*	6-311G*	6-311G**	6-311G**	6-31G*	
R _{CH}	1.094	1.091	1.0908	1.0866	1.085	1.0842	1.0846	1.0846	1.08
R _{γδ}	1.3224	1.319	1.3177	1.3122	1.308	1.3079	1.3199	1.3199	1.29
R _{βγ}	1.2868	1.281	1.2810	1.2789	1.2738	1.2736	1.2846	1.2848	1.29
R _{αβ}	1.2897	1.284	1.285	1.2815	1.2778	1.2777	1.2962	1.2965	1.30
R _{CO}	1.1915	1.181	1.1816	1.1767	1.1668	1.1669	1.1716	1.1716	1.16
A _{HCH}	117.5	117.7	117.9	117.9	117.9	118.2	119.4	119.4	121
A _{HCC}	121.2	121.1	121.1	121.1	121.1	120.9	120.3	119.4	
A _γ	179.9	179.4	179.4	179.9	179.2	173.6	177.2	177.2	
A _β	179.6	176.0	176.0	179.6	174.1	176.5	156.2	155.9	
A _α	179.8	178.7	178.7	179.8	178.3	178.9	174.0	173.9	
D _{HCCC}	90.035	90.200	89.999	90.201	90.019	90.026	90.115	90.650	
v ₁₀ (A')	111(1.7) [†]	45(3.3)	52(3.1)	105(2.2)	67(3.7)	48(3.8)	77(3.9)	91(3.6)	
v ₁₅ (A'')	155(0.2)	136(0.1)	135(0.1)	155(0.01)	137(0.1)	136(0.1)	178(0.6)	193(1.9)	
v ₁₄ (A'')	402(16.8)	389(0.5)	386(0.4)	426(0.7)	403(0.3)	400(0.3)	393(0.2)	403(0.2)	
v ₉ (A')	416(1.0)	395(9.0)	404(8.8)	447(23.9)	450(10.4)	440(9.5)	468(8.5)	466(15.8)	
v ₈ (A')	511(0.4)	478(6.1)	480(6.3)	525(0.1)	510(12.6)	511(12.5)	521(14.8)	502(10.4)	
v ₁₃ (A'')	598(15.9)	600(18.0)	599(18.2)	629(23.2)	634(25.2)	632(25.4)	639(19.3)	629(19.7)	
v ₇ (A')	673(79.2)	665(82.1)	680(77.5)	734(84.7)	723(86.7)	737(81.7)	682(85.6)	729(93.6)	728
v ₆ (A')	744(2.3)	740(2.2)	740(2.2)	769(2.5)	765(2.3)	763(2.3)	777(1.5)	805(2.6)	
v ₁₂ (A'')	962(3.6)	951(3.9)	938(3.3)	990(4.1)	979(4.4)	966(3.8)	964(2.5)	1000(3.8)	
v ₅ (A')	1347(3.0)	1330(2.1)	1321(1.1)	1384(2.5)	1367(1.7)	1358(0.9)	1358(2.2)	1385(3.0)	
v ₄ (A')	1487(84.3)	1474(100.7)	1472(104.5)	1544(82.7)	1531(95.5)	1528(100.7)	1497(45.0)	1517(29.9)	1456
v ₃ (A')	2008(0.03)	1998(0.8)	1999(1.0)	2084(0.7)	2073(5.2)	2073(4.9)	2053(46.3)	2064(37.2)	1996
v ₂ (A')	2290(1503.6)	2276(1685.0)	2276(1696.8)	2359(1975.4)	2344(2173.1)	2345(2193.1)	2326(1725.9)	2312(1506.0)	2242
v ₁ (A')	3083(9.9)	3056(8.4)	3060(12.9)	3167(12.5)	3137(10.7)	3137(16.1)	3186(21.9)	3216(21.2)	3035
v ₁₁ (A'')	3152(0.2)	3126(0.1)	3132(0.8)	3242(0.6)	3213(0.4)	3215(1.6)	3284(2.1)	3306(1.5)	3105

[†]IR intensities are in parentheses. ^aRef. 18. ^bStructure data from Ref. 16; frequency data from Ref. 20.

Table 5. Structures (bond lengths in Å and bond angles in degrees) and harmonic frequencies (cm⁻¹) of H₂C₅O

Method	BLYP	BLYP	BLYP	B3LYP	B3LYP	B3LYP	MP2	MP2 ^a	Exp ^b
Basis set	6-31G*	6-311G*	6-311G**	6-31G*	6-311G*	6-311G**	6-311G**	6-31G*	
R _{CH}	1.0991	1.0966	1.0955	1.0909	1.0890	1.0884	1.0890	1.0891	
R _{δe}	1.3309	1.3265	1.3268	1.3200	1.3166	1.3164	1.3292	1.3291	
R _{γδ}	1.2817	1.2760	1.2761	1.2737	1.2685	1.2685	1.2768	1.2769	
R _{βγ}	1.2971	1.2924	1.2926	1.2887	1.2842	1.2844	1.3002	1.2999	
R _{αβ}	1.2866	1.2823	1.2827	1.2787	1.2750	1.2752	1.3003	1.3000	
R _{CO}	1.1951	1.1852	1.1850	1.1799	1.1702	1.1702	1.1736	1.1736	
A _{HCH}	115.653	115.792	116.101	115.981	116.181	116.445	117.149	117.10	
A _{HCC}	122.173	122.097	121.937	122.022	121.904	121.770	121.376	121.39	
A _{HCC}	122.174	122.111	121.962	121.997	121.915	121.785	121.476	121.510	
A _δ	179.838	179.219	179.143	179.876	179.229	179.075	174.452	174.45	
A _γ	179.890	176.281	178.981	179.985	179.146	178.986	174.275	174.28	
A _β	179.071	176.840	174.424	179.982	174.897	173.916	150.049	150.35	
A _α	179.728	178.893	178.037	179.994	178.426	178.098	172.094	172.13	
v ₁₃ (A')	79(1.4) [†]	27(3.1)	49(3.0)	73(1.8)	53(3.2)	59(3.1)	76(4.0)	83(3.4)	
v ₁₈ (A'')	114(0.0)	102(0.2)	104(0.1)	115(0.01)	106(0.2)	106(0.1)	164(1.2)	170(2.0)	
v ₁₂ (A')	209(1.3)	172(2.2)	171(2.4)	222(1.3)	196(2.1)	197(2.2)	170(2.8)	189(2.3)	
v ₁₇ (A'')	301(2.2)	265(2.1)	265(2.1)	301(2.3)	269(2.1)	269(2.1)	264(2.8)	266(2.4)	
v ₁₁ (A')	390(16.6)	351(14.1)	363(14.1)	428(25.1)	408(18.3)	412(17.5)	396(13.1)	393(14.4)	
v ₁₀ (A')	490(0.9)	457(5.0)	457(6.0)	503(0.2)	495(10.6)	495(11.6)	507(8.8)	485(9.0)	
v ₁₆ (A'')	566(7.4)	566(2.1)	568(1.0)	597(9.4)	595(0.7)	594(0.3)	555(2.6)	566(1.9)	

Table 5. Continued

Method	BLYP	BLYP	BLYP	B3LYP	B3LYP	B3LYP	MP2	MP2 ^a	Exp ^b
Basis set	6-31G*	6-311G*	6-311G**	6-31G*	6-311G*	6-311G**	6-311G**	6-31G*	
v ₁₅ (A'')	630(3.1)	617(2.5)	612(19.0)	651(3.0)	645.7(28.1)	642(28.4)	626(15.6)	612(18.1)	
v ₉ (A')	689(7.5)	626(3.1)	627(3.1)	697(13.4)	646.3(3.0)	647(2.9)	678(3.7)	699(3.9)	
v ₁₄ (A'')	905(31.3)	901(34.5)	909(32.1)	952(33.5)	947(35.8)	955(33.7)	919(34.4)	945(38.4)	
v ₈ (A')	1026(0.1)	1020(0.1)	1010(0.1)	1052(0.1)	1048(0.1)	1038(0.2)	1038(1.0)	1062(0.3)	
v ₇ (A')	1214(22.6)	1204(25.0)	1203(24.8)	1256(23.4)	1244(24.7)	1243(24.5)	1212(4.4)	1228(3.5)	
v ₆ (A')	1468(0.2)	1451(0.02)	1439(0.1)	1506(0.01)	1488(0.3)	1477(0.6)	1483(0.3)	1512(1.2)	
v ₅ (A')	1745(61.4)	1736(60.0)	1734(56.5)	1814(52.2)	1804(48.5)	1802(45.8)	1756(15.3)	1768(17.8)	1726
v ₄ (A')	2118(2.0)	2106(31.1)	2105(3.1)	2193(0.004)	2180(0.7)	2180(1.0)	2154(38.4)	2174(22.7)	2068
v ₃ (A')	2268(1729.4)	2255(1901.2)	2253(1887.5)	2340(2348.5)	2326(2541.6)	2325(2524.8)	2285(1872.3)	2280(1742.5)	2207
v ₂ (A')	3020(23.4)	2098(29.0)	3001(21.6)	3115(15.2)	3089(19.4)	3086(13.1)	3135(10.8)	3166(9.4)	
v ₁ (A')	3080(11.1)	3059(12.0)	3063(8.4)	3183(6.6)	3159(7.1)	3156(4.4)	3223(3.1)	3248(3.9)	

[†]IR intensities are in parentheses. ^aRef. 18. ^bFrequency data ; Ref. 19.

Table 6. Structures (bond lengths in Å and bond angles in degrees) and harmonic frequencies (cm⁻¹) of H₂C₆O

Method	BLYP	BLYP	BLYP	B3LYP	B3LYP	B3LYP	MP2
Basis Set	6-31G*	6-311G*	6-311G**	6-31G*	6-311G*	6-311G**	6-311G**
R _{CH}	1.0947	1.0928	1.0916	1.0874	1.0858	1.0850	1.0859
R _{εζ}	1.3251	1.3211	1.3208	1.3146	1.3108	1.3104	1.3247
R _{δε}	1.2864	1.2810	1.2810	1.2790	1.2739	1.2738	1.2812
R _{γδ}	1.2902	1.2851	1.2851	1.2814	1.2765	1.2765	1.2872
R _{βγ}	1.2898	1.2842	1.2842	1.2813	1.2760	1.2759	1.2872
R _{αβ}	1.2885	1.2841	1.2841	1.2809	1.2766	1.2766	1.2900
R _{CO}	1.1923	1.1825	1.1825	1.1772	1.1673	1.1673	1.1731
A _{HCH}	116.975	117.200	117.515	117.486	117.6644	117.909	119.041
A _{HCC}	121.512	121.400	121.243	121.257	121.1677	121.046	120.479
A _ε	179.979	179.986	179.994	179.973	179.9738	179.972	179.168
A _δ	179.947	179.977	179.966	179.972	179.9523	179.944	175.208
A _γ	179.995	179.873	179.967	179.981	179.8350	179.844	176.453
A _β	179.999	178.928	179.996	179.927	178.6346	178.613	166.457
A _α	179.792	179.661	179.956	179.992	179.6142	179.613	176.846
D _{HCCC}	90.163	90.016	90.046	90.183	89.859	90.011	90.225
v ₂₁ (A'')	74(0.2) [†]	56(0.6)	53(0.1)	88(1.3)	40(1.2)	34(0.8)	
v ₁₅ (A')	82(0.0)	75(0.1)	75(0.1)	109(0.2)	75(0.1)	75(0.6)	
v ₂₀ (A'')	196(0.1)	159(0.04)	157(0.04)	178(0.001)	156(0.01)	155(18.7)	
v ₁₄ (A')	220(1.5)	193(0.9)	191(0.9)	218(0.7)	192(0.8)	190(1.3)	
v ₁₃ (A')	403(0.003)	375(0.6)	368(0.4)	402(0.8)	388(0.7)	381(0.1)	
v ₁₉ (A'')	415(13.4)	412(12.4)	407(14.0)	429(15.3)	424(16.9)	415(0.01)	
v ₁₈ (A'')	489(0.2)	463(5.3)	464(4.4)	519(11.1)	499(10.6)	498(9.5)	
v ₁₂ (A')	547(1.5)	543(1.5)	543(1.4)	552(1.9)	560(1.6)	548(0.05)	
v ₁₇ (A'')	552(10.9)	548(1.0)	545(0.3)	571(0.0006)	566(0.2)	560(1.5)	
v ₁₁ (A')	647(1.1)	563(9.8)	560(7.0)	611(5.5)	596(9.8)	588(2.1)	
v ₁₀ (A')	696(6.2)	622(10.0)	603(13.2)	639(22.7)	635(19.2)	617(27.7)	
v ₁₆ (A'')	716(74.3)	708(73.3)	720(69.1)	778(75.4)	768(76.0)	781(72.0)	
v ₉ (A')	966(1.9)	958(2.1)	946(1.7)	992(2.1)	988(2.1)	975(1.7)	
v ₈ (A')	1068(39.0)	1059(45.1)	1059(45.5)	1078(48.0)	1094(43.8)	1094(44.1)	
v ₇ (A')	1391(9.7)	1372(12.7)	1361(13.8)	1409(11.1)	1409(10.3)	1399(11.7)	
v ₆ (A')	1567(0.5)	1557(0.1)	1556(12.5)	1591(1.2)	1616(0.04)	1616(0.01)	
v ₅ (A')	1941(255.2)	1932(302.2)	1932(310.3)	1966(358.3)	2003(310.0)	2003(320.6)	
v ₄ (A')	2164(368.3)	2151(465.7)	2151(477.7)	2179(856.2)	2217(943.8)	2218(968.6)	
v ₃ (A')	2271(2580.5)	2260(2795.0)	2261(2802.7)	2277(3516.0)	2329(3558.6)	2329(3562.2)	
v ₂ (A')	3073(5.2)	3046(3.9)	3049(8.3)	3114(5.6)	3128(5.4)	3127(10.9)	
v ₁ (A')	3141(0.01)	3115(0.03)	3120(6.4)	3189(0.1)	3204(0.05)	3204(0.7)	

[†]IR intensities are in parentheses (KM/Mole)

Table 7. Structures (bond lengths in Å and bond angles in degrees) and harmonic frequencies (cm^{-1}) of H_2C_7O

Method	BLYP	BLYP	BLYP	B3LYP	B3LYP	B3LYP	MP2
Basis set	6-31G*	6-311G*	6-311G**	6-31G*	6-311G*	6-311G**	6-311G**
R_{CH}	1.0981	1.0957	1.0947	1.0900	1.0882	1.0875	1.0889
$R_{C\eta}$	1.3312	1.3271	1.3269	1.3201	1.3165	1.3165	1.3303
$R_{e\zeta}$	1.2814	1.2758	1.2758	1.2734	1.2681	1.268	1.2761
$R_{\delta e}$	1.2985	1.2935	1.2936	1.2901	1.2854	1.2858	1.2954
$R_{\gamma\delta}$	1.2846	1.2791	1.2791	1.276	1.2709	1.2709	1.2807
$R_{\beta\gamma}$	1.2943	1.2889	1.2889	1.2856	1.2805	1.2808	1.2908
$R_{\alpha\beta}$	1.2877	1.2830	1.2830	1.2798	1.2757	1.2758	1.2929
R_{CO}	1.1939	1.1842	1.1842	1.1786	1.1689	1.1690	1.1741
A_{HCH}	115.986	116.146	116.426	116.344	116.499	116.777	117.431
A_{HCC}	122.018	121.928	121.794	121.830	121.758	121.635	121.307
A_{HCC}	121.996	121.926	121.780	121.826	121.743	121.588	121.262
A_{ζ}	179.970	179.932	179.802	179.928	179.858	179.854	177.993
A_{ϵ}	179.946	179.974	179.907	179.952	179.946	179.926	178.107
A_{δ}	179.972	179.976	179.960	179.972	179.762	179.522	173.322
A_{γ}	179.998	179.999	179.957	179.982	179.507	179.495	174.434
A_{β}	179.956	179.932	179.871	179.901	179.053	176.188	159.336
A_{α}	179.970	179.970	179.929	179.899	179.098	178.803	175.089
$\nu_{17}(A')$	57(0.3) [†]	35(0.8)	32(0.8)	54(0.4)	44(1.0)	49(0.9)	
$\nu_{24}(A'')$	64(0.02)	60(0.2)	60(0.2)	65(0.03)	61(0.2)	61(0.2)	
$\nu_{16}(A')$	142(0.5)	111(0.2)	110(0.2)	136(0.3)	121(0.2)	126(0.2)	
$\nu_{23}(A'')$	176(1.6)	155(1.4)	155(1.5)	177(1.7)	157(1.4)	158(1.4)	
$\nu_{15}(A')$	255(2.9)	224(5.2)	220(5.2)	268(3.7)	247(5.5)	246(5.1)	
$\nu_{22}(A'')$	330(0.001)	300(1.0)	299(0.9)	329(0.03)	303(1.3)	302(1.3)	
$\nu_{14}(A')$	399(12.3)	383(6.4)	374(6.0)	440(18.6)	425(6.5)	429(3.8)	
$\nu_{13}(A')$	476(0.2)	457(7.7)	453(8.5)	488(1.4)	493(5.0)	492(5.6)	
$\nu_{12}(A')$	483(1.8)	479(1.8)	479(1.7)	498(1.8)	497(13.9)	500(16.4)	
$\nu_{21}(A'')$	538(10.2)	536(6.8)	536(6.0)	570(12.4)	559(4.1)	554(5.0)	
$\nu_{20}(A'')$	610(1.6)	578(7.7)	573(1.5)	624(4.9)	593(4.6)	555(3.2)	
$\nu_{11}(A')$	655(2.6)	612(2.5)	577(8.2)	653(3.5)	611(16.5)	611(16.3)	
$\nu_{19}(A'')$	780(3.6)	724(6.5)	732(6.6)	780(6.7)	736(11.5)	724(12.0)	
$\nu_{18}(A'')$	873(37.5)	868(40.8)	877(37.5)	921(40.3)	915(41.9)	923(39.3)	
$\nu_{10}(A')$	946(24.4)	939(27.7)	939(28.0)	978(25.3)	970(27.9)	970(28.1)	
$\nu_9(A')$	1013(0.002)	1008(0.002)	998(0.02)	1040(0.01)	1035(0.01)	1024(0.04)	
$\nu_8(A')$	1354(5.7)	1343(6.3)	1340(5.6)	1396(3.3)	1383(3.9)	1379(3.4)	
$\nu_7(A')$	1475(28.2)	1458(22.9)	1450(21.6)	1518(26.1)	1502(20.3)	1494(18.3)	
$\nu_6(A')$	1753(42.5)	1744(52.5)	1744(54.9)	1819(43.9)	1810(51.9)	1808(53.6)	
$\nu_5(A')$	2054(142.1)	2045(152.3)	2044(146.9)	2127(121.6)	2117(129.6)	2116(121.5)	
$\nu_4(A')$	2141(539.8)	2128(593.0)	2128(572.3)	2204(1027.7)	2192(1115.8)	2190(1086.5)	
$\nu_3(A')$	2260(2644.0)	2251(2887.0)	2251(2891.1)	2330(3531.4)	2320(3795.1)	2319(3799.7)	
$\nu_2(A')$	3031(18.6)	3008(23.4)	3010(5.6)	3124(9.7)	3098(13.1)	3096(6.6)	
$\nu_1(A')$	3094(6.9)	3073(7.5)	3075(4.7)	3195(3.6)	3170(3.9)	3169(1.9)	

[†]IR intensities are in parentheses (KM/Mole)

frequencies to within 30 cm^{-1} . The CH_2 symmetric and anti-symmetric stretching frequencies are very high, in the range of $3000-3100 \text{ cm}^{-1}$. The B3LYP method seems to give too large values for high frequency stretching modes, as we observed before.⁵⁻⁹ It can be seen that the floppy modes involving the C_{β} carbon atom is in-plane for H_2C_3O , H_2C_5O , and H_2C_7O , while those for H_2C_4O and H_2C_6O are out of plane (Figure 1). The harmonic frequencies for this floppy mode of H_2C_nO ($n \geq 4$) are predicted to be much smaller

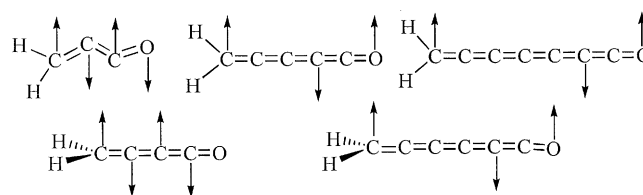


Figure 1. Floppy vibrational modes of cumulenones.

than that of $\text{H}_2\text{C}_3\text{O}$, in accordance with the trend of the kinky angle discussed above, and they stabilize to about 50 cm^{-1} ($\nu = 161, 52, 49, 53, 32\text{ cm}^{-1}$ for $n = 3-7$). Again, agreement of the BLYP results with the experimentally measured (matrix IR)^{19,20} frequencies is very good.

It must be stressed that most of the present computation were performed on a personal computer. Bigger molecules such as $\text{H}_2\text{C}_6\text{O}$ and $\text{H}_2\text{C}_7\text{O}$ presumably cannot be treated using personal computers by employing more conventional correlated *ab initio* methods. Our present work therefore showed that the density functional theory can give results that are comparable to or better than the conventional methods (except CCSD or CCSD(T)) in much more economical fashion.

Conclusion

We presented the density functional theory calculations on cumulenones $\text{H}_2\text{C}_n\text{O}$ ($n = 3-7$). Good agreement between the computed properties and the experimental observations indicates the efficiency of the BLYP method for medium-sized carbon clusters. We hope that the present calculations would stimulate further investigations on these interesting molecules.

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