Ruthenium-NHC Complexes Immobilized on MCF: Application to Catalytic Ring-Closing Metathesis

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Diverse organometallic systems are used as useful catalysts in countless organic reactions, including the C-C bond formation. Since the discovery of olefin metathesis, it has been recognized as a powerful synthetic tool for the preparation of diverse chemicals.² For example, a ring-closing metathesis (RCM) catalyzed by Grubbs catalysts has been used for construction of diverse cyclic organic compounds.³ The series of Grubbs catalysts have many advantages as catalytic compounds such as facile control of their chemical properties. However, to be applied as industrial catalysts, the catalysts should be heterogenized. In literature, various ways to immobilize Grubbs catalysts have been suggested.⁴ Some groups used the polymers such as butyldiethylsily polystyrene (PS-DES) or polyethylene glycol (PEG) as a support material (Figure 1A). 4a There is also the case that functionalized monolithic sol-gel was used.4b In addition, imidazolium styrenyl ether ligand was used and immobilized to ionic liquid-support (Figure 1B).4c Recently, fluorous ligand was used to develop fluorous catalysts and through fluorous solid-phase extraction, a heterogeneous catalyst was developed (Figure 1C).4e In the existing researches, it was difficult to synthesize catalysts or it was restricted to synthesize various catalysts. Therefore, these researchers thought it would be necessary to develop a catalyst which can be easily synthesized and is robust and stable enough to be recycled without any problem.

We herein report on the development of an excellent heterogeneous catalyst for RCM by immobilizing a modified Grubbs catalyst on mesoporous silica. Mesoporous and structurally uniform materials, including MCM, SBA, KIT and MCF, have been applied to the diverse heterogeneous catalytic systems.⁵ In this study, siliceous mesostructured cellular foam (MCF) ⁶ was chosen as a solid support, which

have a large surface area of $700-1005 \text{ m}^2/\text{g}$ and a pore volume within the range of $1.4-2.3 \text{ cm}^3/\text{g}$.

Recently, *N*-heterocyclic carbenes (NHCs) have emerged as promising ligands for the design of new homogeneous catalysts. In particular, the NHC-based Grubbs catalysts have attracted much attention because their ease of preparation and their chemical and topological versatility allows the preparation of a wide variety of complexes whose chemical properties can be easily modulated. Thus, we decided to immobilize ruthenium-NHC catalysts on mesoporous silica. Mono-substituted imidazoles and benzimidazole were synthesized by a known procedure (Eq. 1).

$$(iii) \longrightarrow (iv) \longrightarrow (Si \longrightarrow N \longrightarrow N \longrightarrow R$$

$$Cl \longrightarrow Ru \longrightarrow Cl \longrightarrow PCy_3 Ph$$

$$1-5 \longrightarrow PCy_3 Ph$$

Reagents and conditions: (i) 60 °C, overnight; (ii) MCF, CH₂Cl₂, N₂, reflux, 12 h; (iii) K'OBu, THF, 1 h.; (iv) Grubbs' catalyst, 12 h.

Figure 1. Examples of Immobilized Grubbs catalysts.

Figure 2. Catalysts used in this work.

The synthesized monosubstituted imidazoles were reacted with (3-bromopropyl) trimethoxysilane to make precursor **1a-5a**. Compounds **1a-5a** were linked to MCF under the condition of refluxing dichloromethane. Reaction of **1a-5a** with Grubbs' catalyst yielded catalysts 1-5. The synthesized catalysts were analyzed by combustion analysis and ICS-AES (Eq. 2).¹⁷

Using *N*,*N*-diallyl-4-methylbenzenesulfonamide as a model substrate, we studied the performance of catalysts in the RCM reaction (Eq. 3 and Table 1).

TsN
$$CH_2Cl_2, r.t., 3h$$
 TsN (3)

As shown in the Table 1, the best result (93%) was obtained with catalyst 1, and this catalyst was reused at least seven times without any significant loss of catalytic activity. An inductively coupled plasma-mass spectrometry (ICP-MS) study showed that the ruthenium bled from the catalyst was negligible, ranging from 0.00074% to 0.00519%. These results confirm that the catalytic system presented here satisfies the conditions for heterogeneous catalysts of easy separation, recyclability, and persistence.

Next, we investigated the versatility of dienes (Table 2). High yields were observed for all dienes except entries 3 and 11. In particular, the use of a diene with a substituted double bond (entry 11) diminished the yield to 30%. However, no reaction was observed for an intermolecular RCM (entry 12). We could get one with satisfactory yield not influenced by the functional group of various dienes. However, in case there is a hydroxyl group, the yield went down to 53%, which seems to have resulted from that substrate is absorbed by hydrophilic silica.

As shown in Table 1, the catalyst had no problem with its activity when it was re-used (no. 7). However, our catalyst system uses silica as a supporting body, but silica disintegrates after recycling, a problem that is more serious than

Table 1. Ring closing metathesis of *N*,*N*-diallyl-4-methylbenzene-sulfonamide by the supported catalysts^a

| Entry | Catalyst | Yield (%) | Ru Residue (ppm) ^b | |
|-------|--------------------|-----------|----------------------------------|--|
| 1 | 1 | 93 | 1.2 | |
| 2 | recovered from #1 | 91 | 1.3 | |
| 3 | recovered from #2 | 89 | 2.1 | |
| 4 | recovered from #3 | 88 | 0.3 | |
| 5 | recovered from #4 | 88 | 0.3 | |
| 6 | recovered from # 5 | 89 | 0.4 | |
| 7 | recovered from #6 | 90 | 0.4 | |
| 8 | 2 | 90 | | |
| 9 | 3 | 79 | | |
| 10 | 4 | 73 | | |
| 11 | 5 | 85 | | |

^aIsolated yield. ^bDetermined by ICP-Mass spectroscopic analysis

ruthenium metal leaching. Generally, this problem became serious when the catalytic system was used more than three times. Separation of the catalysts from the reaction mixture became tough and annoying. To overcome this problem, we introduced a flow micro reactor (Figure 2) and used it for the RCM reaction. However, when the reaction was carried out in the flow micro reactor, the reaction gave 50% yield in

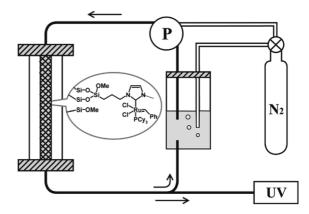


Figure 3. Flow microreactor used for RCM.

Table 2. Olefin metathesis using 1

| entry | substrate | product | yield (%) ^a | entry | substrate | product | yield (%) ^a |
|-------|--------------------|--------------------|------------------------|-------|------------------------|------------------------|------------------------|
| 1 | TsN | TsN | 93 | 7 | (t-Bu)O ₂ C | (f-Bu)O ₂ C | 88 |
| 2 | MeO ₂ C | MeO ₂ C | 90 | 8 | BnO | BnO | 73 |
| 3 | но | но | 53 | 9 | | | 91 |
| 4 | MeO MeO | MeO MeO | 75 | 10 | MeO ₂ C | MeO ₂ C | 79 |
| 5 | | | 92 | 11 | TsN | TsN | 30 |
| 6 | PhO ₂ | PhO ₂ | 87 | 12 | MeO ₂ C + | OMe | - |

^aIsolated yield.

case of *N*,*N*-diallyl-4-methylbenzenesulfonamide for 1 hour. We expected that lengthening a reaction time would increase a yield. However, because of the restriction in the flow micro reactor system, it was impossible to get a result for a longer reaction time. This problem will be overcome by developing a more reactive catalyst.

In summary, we developed heterogeneous Ru-NHC catalysts by immobilization of Grubbs catalysts on the MCF-type mesoporous silica through a modification of the NHC ligand. Starting from NHC ligand, a facile three-step reaction gave the supported catalyst which is a robust, stable, and easy to reuse. In addition to the advantages discussed above, the developed catalyst was highly effective for the intramolecular RCM reaction and could be reused at least seven times without any loss of its catalytic activities. Moreover, research is being conducted to further develop our RCM catalysts system for application to enantioselective catalysis.

Experimental Section

Representative procedure of the catalytic ring-closing metathesis (RCM) reaction. To a flame-dried 25 mL Schlenk flask capped with a rubber septum, a diene substrate (0.40 mmol) and NHC-bounded mesoporous silica catalyst (0.2 g, 0.02 mmol Ru) were dissolved in 10 mL of CH₂Cl₂. After the resulting solution was stirred for 3 h, the solution

was filtered, concentrated, and chromatographed on a silica gel column eluting with hexane/diethyl ether (10:1) to give the product. When the catalyst was recycled, the catalyst was filtered, dried in vacuum, and reused for the further catalytic reaction.

Modification of MCF surface with NHC ligands. MCF (2 g) was dispersed in dry CH₂Cl₂ (25 mL) at room temperature under N₂, after which ligand (1a-5a) (0.3 mmol) was added. The mixture was filtered after being stirred for 12 h. After drying in a vacuum for 1 day, NHC-bounded mesoporous silica was obtained.

Representative procedure of the Ruthenium-NHC complexes-immobilized on MCF. To a flame-dried 25 mL Schlenk flask capped with a rubber septum, NHC-bounded mesoporous silica (silica were combined with 0.3 mmol of NHC ligand) were dissolved in 10 mL of THF and K'OBu (0.4 mmol) was added. After the mixture was stirred for 1h, Grubbs' catalyst (0.3 mmol) was added. The mixture was filtered after being stirred for 12 h and the brown product was washed with 50 mL of pentane. After drying in a vacuum for 1 day, the Ruthenium-NHC complexes-immobilized on MCF was obtained.

1-Methyl-3-(3-trimethoxysilanopropyl)-imidazolium (1a): ¹H NMR (CDCl₃, 300 MHz): δ 0.67 (d, J = 8.0 Hz, 2H), 2.04 (m, 2H), 3.58 (s, 9H), 4.13 (s, 3H), 4.35 (t, J = 7.8 Hz, 2H), 7.47 (s, 1H), 7.59 (s, 1H), 9.84 (s, 1H); ¹³C NMR (CDCl₃, 75 MHz): δ 6.1, 24.3, 37.3, 51.0, 52.0, 122.3, 123.8, 137.14; exact mass for (C₁₀H₂₀SiO₃N₂): [M+1] calcd 244.1243, obsd 245.1322.

1-Phenyl-3-(3-trimethoxysilanopropyl)-imidazolium (*2a*): 1 H NMR (CDCl₃, 300 MHz): δ 0.73 (t, J= 7.7 Hz, 2H), 2.12 (m, 2H), 3.59 (s, 9H), 4.61 (t, J= 7.2 Hz, 2H), 7.58 (m, 5H), 7.81 (d, J= 6.8 Hz, 2H), 10.56 (s, 1H); 13 C NMR (CDCl₃, 75 MHz): δ 0.2, 6.1, 24.4, 50.8, 51.0, 52.5, 120.8, 122.2, 123.3, 130.7, 130.9, 134.5, 135.6; exact mass for (C₁₅H₂₂SiO₃N₂): [M+1] calcd 306.1400, obsd 307.1480.

1-(4-t-Butyl-phenyl)-3-(3-trimethoxysilanopropyl)-imidazolium (3a): 1 H NMR (CDCl₃, 300 MHz): δ 0.73 (t, J = 7.9 Hz, 1H), 1.34 (s, 9H), 2.12 (m, 2H), 3.58 (s, 9H), 4.60 (t, J = 7.1 Hz, 2H), 7.59 (d, J = 6.7 Hz, 2H), 7.72 (d, J = 6.5 Hz, 4H), 10.46 (s, 1H); 13 C NMR (CDCl₃, 75 MHz): δ 6.0, 24.3, 31.3, 35.1, 50.7, 51.0, 52.3, 121.0, 121.7, 123.4, 126.0, 127.7, 132.0, 135.1, 154.1; exact mass for (C₁₉H₃₀SiO₃N₂): [M+1] calcd 362.2026, obsd 363.2106.

1-Methyl-3-(3-trimethoxysilanopropyl)-4,5-diphenyl-imidazolium (4a): 1 H NMR (CDCl₃, 300 MHz): δ 0.63 (t, J = 6.8 Hz, 2H), 1.91 (m, 2H), 3.50 (s, 9H), 4.23 (t, J = 8.1 Hz, 2H), 5.31 (s, 3H), 7.41 (m, 10H), 10.32 (s, 1H); 13 C NMR (CDCl₃, 75 MHz): δ 6.2, 10.6, 24.0, 27.4, 35.0, 35.5, 49.9, 50.7, 50.9, 124.8, 125.1, 126.1, 127.6, 129.3, 129.4, 129.9, 130.5, 130.6, 130.7, 130.8, 131.0, 131.9, 132.7, 136.9; exact mass for (C₁₈H₂₈SiO₃N₂): [M+1] calcd 396.1869, obsd 397.1945.

1-(2,4,6-Trimethyl-phenyl)-3-(3-trimethoxysilanopropyl)imidazolium (5a): 1 H NMR (CDCl₃, 300 MHz): δ 0.71 (t, J = 8.0 Hz, 2H), 2.10 (m, 8H), 2.32 (s, 3H), 3.59 (s, 9H), 4.71 (t, J = 7.1 Hz, 2H), 7.02 (s, 2H), 7.25 (s, 1H), 7.79 (s, 1H), 9.91 (s, 1H); 13 C NMR (CDCl₃, 75 MHz): δ 10.7, 18.0, 21.3, 27.4, 50.8, 51.0, 126.0, 130.1, 134.4, 141.6; exact mass for (C₁₈H₂₈SiO₃N₂): [M+1] calcd 348.1869, obsd 349.1941.

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References

- (a) Ritleng, V.; Sirlin, C.; Pfeffer, M. Chem. Rev. 2002, 102, 1731.
 (b) Terao, J.; Kambe, N. Bull. Chem. Soc. Jpn. 2006, 79, 663. (c) Metal-Catalyzed Cross-Coupling Reactions, 2nd ed.; de Meijere, A., Diederich, F., Eds.; Wiley-VCH: Weinheim, 2004.
- (a) Mol, J. C. J. Mol. Catal. A 2004, 213, 39. (b) Hong, J. H.; Ko, O. H. Bull. Korean Chem. Soc. 2003, 24, 1289.
- 3. (a) Applications in Organic Synthesis in Handbook of Metathesis; Grubbs, R. H., Ed.; Wiley-VCH: Weinheim, 2003; Vol. 2. (b) Chatterjee, A. K.; Choi, T.-L.; Sanders, D. P.; Grubbs, R. H. J. Am. Chem. Soc. 2003, 125, 11360. (c) Bielawsky, C. W.; Grubbs, R. H. Angew. Chem. Int. Ed. 2000, 39, 2903.
- 4. (a) Yao, Q. Angew. Chem. Int. Ed. 2000, 39, 3896. (b) Kingsbury,

- J. S.; Garber, S. B.; Giftos, J. M.; Gray, B. L.; Okamoto, M. M.; Farrer, R. A.; Fourkas, J. T.; Hoveyda, A. H. *Angew. Chem. Int. Ed.* **2001**, *40*, 4251. (c) Audic, N.; Clavier, H.; Mauduit, M.; Guillemin, J.-C. *J. Am. Chem. Soc.* **2003**, *125*, 9248. (d) Buchmeiser, M. R. *New J. Chem.* **2004**, *28*, 549. (e) Matsugi, M.; Curran, D. P. *J. Org. Chem.* **2005**, *70*, 1636. (f) Varray, S.; Lazaro, R.; Martinez, J.; Lamaty, F. *Organometallics* **2003**, *22*, 2426. (g) Kingsbury, J. S.; Harrity, J. P. A.; Bonitatebus, P. J.; Hoveyda, A. H. *J. Am. Chem. Soc.* **1999**, *121*, 791. (h) Grela, K.; Tryznowski, M.; Bieniek, M. *Tetrahedron Lett.* **2002**, *43*, 9055.
- Ko, K.; Park, S.; Choi, M. Bull. Kor. Chem. Soc. 2000, 21, 951. (b)
 Kim, S.-W.; Son, S. U.; Lee, S. I.; Hyeon, T.; Chung, Y. K. J. Am. Chem. Soc. 2000, 122, 1550.
- Schmidt-Winkel, P.; Lukens, W. W., Jr.; Yang, P.; Margolese, D. I.; Lettow, J. S.; Ying, J. Y.; Stucky, G. D. *Chem. Mater.* 2000, 12, 686.
- (a) Wanzlick, H.-W. Angew. Chem. 1962, 74, 129. (b) Wanzlick, H.-W.; Esser, F.; Kleiner, H. J. Chem. Ber. 1963, 96, 1208. (c) Wanzlick, H.-W.; Kleiner, H. J. Chem. Ber. 1963, 96, 3024.
- Arduengo, A. J., III; Harlow, R. L.; Kline, M. J. Am. Chem. Soc. 1991, 113, 361.
- (a) Trnka, T. M.; Grubbs, R. H. Acc. Chem. Res. 2001, 34, 18. (b) Handbook of Metathesis; Grubbs, R. H., Ed.; Wiley-VCH: Weinheim, Germany, 2003; Vols 1-3. (c) Blechert, S.; Connon, S. J. Angew. Chem., Int. Ed. 2003, 42, 1900. (d) Schrock, R. R.; Hoveyda, A. H. Angew. Chem., Int. Ed. 2003, 42, 4592.
- (a) Herrmann, W. A.; Öfele, K.; Preysing, D. V.; Schneider, S. K. J. Organomet. Chem. 2003, 687, 229. (b) Navarro, O.; Kelly, R. A., III; Nolan, S. P. J. Am. Chem. Soc. 2003, 125, 16194. (c) Marion, N.; Ecarnot, E. C.; Navarro, O.; Amoroso, D.; Bell, A.; Nolan, S. P. J. Org. Chem. 2006, 71, 3816. (d) Marion, N.; Navarro, O.; Mei, J.; Stevens, E. D.; Scott, N. M.; Nolan, S. P. J. Am. Chem. Soc. 2006, 128, 4101. (e) Navarro, O.; Marion, N.; Mei, J.; Nolan, S. P. Chem.-Eur. J. 2006, 12, 5142. (f) Jung, I. G.; Seo, J.; Lee, S. I.; Choi, S. Y.; Chung, Y. K. Organometallics 2006, 25, 4240.
- De Lewis, A. K.; Caddick, S.; Cloke, F. G. N.; Billingham, N. C.; Hitchcock, P. B.; Leonard, J. J. Am. Chem. Soc. 2003, 125, 10066.
- (a) Lee, H. M.; Jiang, T.; Stevens, E. D.; Nolan, S. P. *Organometallics* **2001**, *20*, 1255. (b) Hillier, A. C.; Lee, H. M.; Stevens, E. D.; Nolan, S. P. *Organometallics* **2001**, *20*, 4246. (c) Vasquez-Serrano, L. D.; Owens, B. T.; Buriak, J. M. *Chem. Commun.* **2002**, 2518. (d) Perry, M. C.; Cui, X.; Powell, M. T.; Hou, D.-R.; Reibenspies, H. J.; Burgess, K. *J. Am. Chem. Soc.* **2003**, *125*, 113.
- (a) Markò, I. E.; Stérin, S.; Buisine, O.; Mignani, G.; Branlard, P.; Tinant, B.; Declerq, J.-P. *Science* 2002, 298, 204. (b) Díez-González, S.; Kaur, H.; Kauer Zinn, F.; Stevens, E. D.; Nolan, S. P. *J. Org. Chem.* 2005, 70, 4784.
- 14. Zuo, G.; Louie, J. Angew. Chem., Int. Ed. 2004, 43, 2277.
- (a) Öfele, K.; Herrmann, W. A.; Mihalios, D.; Elison, M.; Herdtweck, E.; Scherer, W.; Mink, J. J. Organomet. Chem. 1993, 459, 177. (b) Huang, I.; Schanz, H.-J.; Stevens, E. D.; Nolan, S. P. Organometallics 1999, 18, 2370.
- (a) Bourissou, D.; Guerret, O.; Gabbai, F. P.; Bertrand, G. Chem. Rev. 2000, 100, 39. (b) Díez-González, S.; Nolan, S. P. Coord. Chem. Rev. 2007, 251, 874.
- 17. Through the element analysis, nitrogen was checked, through which the existence of NHC ring could be confirmed indirectly. Amount of Nitrogen in Catalyst 1: 2.7277%, 2: 0.8561%, 3: 1.0663%, 4: 1.3447%, 5: 0.9415%. Showing loading of catalyst (0.11-0.175 mmol Ru/g: ICP-AES analysis).