

Correlation between the Width of Gallium Oxide Nanobelts and the Diameter of the Catalysts

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One-dimensional (1D) nanostructures such as wires, rods, and belts have been the focus of extensive studies owing to their unique physical properties and potential to revolutionize broad areas of nanotechnology.¹ Various synthetic techniques have been used to grow 1D nanostructures *via* the vapor-liquid-solid (VLS) mechanism,² which was proposed by Wagner and Ellis.³ The characteristic features of this mechanism are that the 1D nanostructures have metal or alloy droplets at their tips which act as catalysts and these droplets define their diameters, *i.e.*, smaller catalyst droplets yield thinner 1D nanostructures. Very recently some theoretical models have been suggested on the effect of the droplet size on the growth kinetics of nanowires by the VLS mechanism.⁴ However, the effect of the diameter (d_{cat}) of the catalyst droplets on the diameter (d_{nano}) of the 1D nanostructures remains little understood theoretically.⁵ There have only been a few reports which show an empirical correlation between d_{nano} and d_{cat} .⁵⁻⁷ Another goal of current theoretical investigations is to find out the main factor which determines the ratio of d_{nano} to d_{cat} . For example, in most cases, the d_{nano} value of the nanowires was found to be greater than the initial d_{cat} value of the catalyst prior to their growth.⁷⁻¹⁰ On the other hand, the d_{nano} value of GaAs nanowires, which were formed by the solution-liquid-solid mechanism, was *ca.* 60% of the initial d_{cat} value.⁶

In this communication, we report the synthesis of a monoclinic β -Ga₂O₃ nanobelts *via* the reduction of Ga₂O₃ by CO and subsequent oxidation and show an empirical corre-

lation between the d_{nano} value of the nanobelts and the d_{cat} value. β -Ga₂O₃ powder (Wako Pure Chemical Industries, Ltd., 99.99%) was loaded in an alumina crucible and a (100) silicon substrate was placed at the downstream end of the crucible at a distance of 200-250 mm from the Ga₂O₃ source. The alumina crucible was heated at 1200 °C in a mixed gas flow (90% nitrogen and 10% CO) for 2 h. After cooling the furnace to room temperature, the as-deposited products on the surface of the substrate were wool-like with a light gray color. The XRD pattern of the products revealed that they are assigned to a crystalline monoclinic β -Ga₂O₃ (JCPDS Card No. 41-1103).

The morphology of the β -Ga₂O₃ was influenced by the substrate temperature. The product deposited on the substrate at *ca.* 720 °C is demonstrated in Figure 1(a), in which there are lots of nanobelts having droplets at their tips. As shown in Figure 1(b), a typical nanobelt in the product terminates in a droplet at its tip. Energy-dispersive X-ray spectroscopy measurements made on the droplet and the belt indicated that the droplet is composed mainly of Ga, while the belt is composed of Ga and oxygen. The presence of the Ga droplet at the nanobelt tip in Figure 1 provides strong evidence that the nanobelt grows *via* the VLS mechanism. β -Ga₂O₃ powder in a hot zone will be reduced first to Ga₂O(g) and then to Ga(g) by CO. As gaseous Ga travels in the gas stream to cooler zones in the furnace, it will be oxidized to Ga₂O₃(g) by O₂. The Ga₂O₃(g) diffuses into the unoxidized Ga droplets on the Si substrate, and then is supersaturated

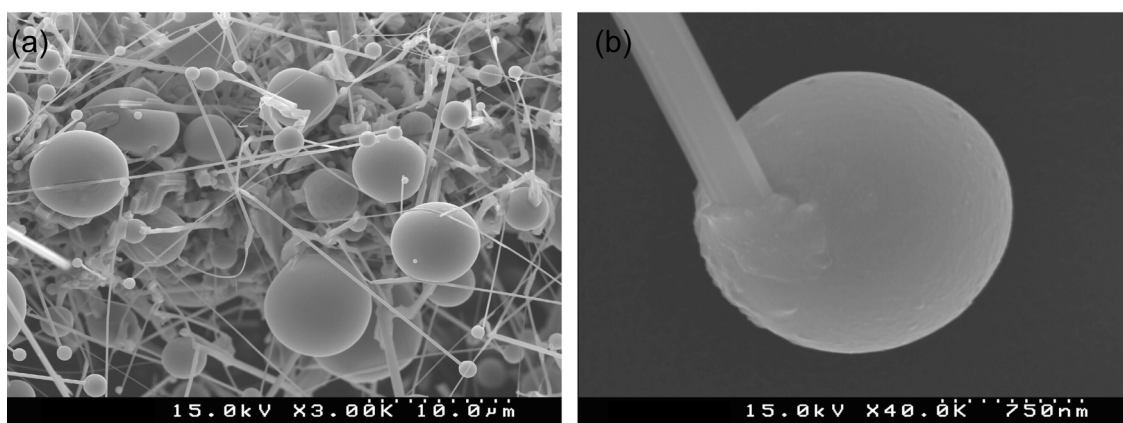


Figure 1. (a) SEM image of the as-synthesized products obtained from the reduction of β -Ga₂O₃ by CO and subsequent oxidation. (b) A typical SEM image of a β -Ga₂O₃ nanobelt carrying a Ga droplet at its tip.

and nucleated in the droplets. Continuous diffusion causes Ga₂O₃ nanobelts to grow out of the droplet, as shown in Figure 1(b). The Ga₂O₃ nanobelts were obtained in a flowing gas mixture of nitrogen and CO without any intentional supply of O₂. Therefore, the residue oxygen in the mixed gas is believed to be the source of oxygen for the growth of the nanostructures. The formation of Ga droplets indicates that the amount of the residue oxygen was not sufficient to oxidize all of the Ga vapor formed. Fu *et al.* synthesized β -Ga₂O₃ nanobelts via the reduction of Ga₂O₃ by H₂ and subsequent oxidation at 1000 °C but did not observe any Ga droplets at the nanobelt tips.¹¹ They surmised that the Ga droplets might be reduced during the growth periods at the high temperature (900 °C) of the deposition site because of the low melting point (29.8 °C) of Ga metal.

As shown in Figure 1(a), the Ga droplets at the nanobelt tip were inhomogeneous and relatively large (above 500 nm). The analysis by SEM revealed that the width (w_{belt}) of the β -Ga₂O₃ nanobelts depends on the d_{cat} value of the Ga droplets. Figure 2 shows that there is a linear correlation between w_{belt} and d_{cat} . The data in Figure 2 were taken from the spherical Ga droplets. The equation of the correlation line was obtained through linear regression analysis and is as follows:

$$w_{belt}(\text{nm}) = (0.21 \pm 0.01) d_{cat}(\text{nm}) + (5.3 \pm 16)$$

where the errors shown correspond to the standard deviations. The slope of the correlation line means that the w_{belt} value is *ca.* 20% of the d_{cat} value. This contrasts with the results obtained for the growth of β -Ga₂O₃ nanobelts by the Sn-catalyzed VLS mechanism, in which the d_{cat} value of the Sn particles was much smaller than the w_{belt} value of the nanobelt and was the same as the thickness of the nanobelt.¹² This difference might be explained in terms of the solubility of the growth vapor in the catalyst droplets. The solubility is expected to increase with increasing temperature gap between the melting point of the catalyst and the growth temperature, *i.e.*, the solubility of Ga₂O₃ vapor will be much higher in the Ga droplets than in the Sn droplets higher at the same temperature, because the melting point of Ga is much lower than that of Sn. This higher solubility will cause the initial size of the catalyst droplets to be increased, *i.e.*, the increase in the size will be larger for the Ga droplets than for the Sn droplets. In addition, the degree of the increase will be a function of the reaction temperature.

Besides the size of the catalyst, the growth temperature also has an effect on the size of the 1D nanostructures. For example, Chang and Wu synthesized β -Ga₂O₃ nanowires on Au-coated Si substrates by evaporating Ga(acetylacetonate)₃ and found that there was no significant distinction between the average size of the Au particles formed at 850 and 550 °C, whereas the diameters of the nanowires increased by a factor of 2 over this temperature range.¹³ Recently we showed that the growth temperature influenced the thickness of aluminum nitride (AlN) whiskers which were grown via the Fe-catalyzed VLS mechanism, *i.e.*, a higher growth

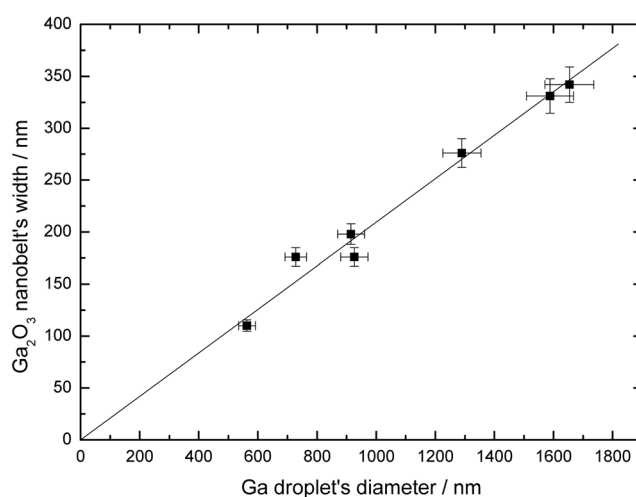


Figure 2. Correlation between the width of the β -Ga₂O₃ nanobelts and the diameter of the Ga droplets.

temperature yielded thicker AlN whiskers.¹⁴ Therefore, the temperature should be considered when analyzing the ratio of d_{nano} (or w_{belt}) to d_{cat} . More detailed studies are in progress in order to elucidate the effects of the catalyst and temperature on the ratio of w_{belt} to d_{cat} in the growth of β -Ga₂O₃ nanobelts.

In conclusion, β -Ga₂O₃ nanobelts were prepared by the thermal evaporation of β -Ga₂O₃ in a flowing gas mixture of nitrogen and CO without using any catalyst. The nanobelts were deposited on a Si substrate *via* the VLS mechanism. The w_{belt} value of the nanobelts was linearly correlated with the d_{cat} value of the Ga droplets, suggesting that the width of the β -Ga₂O₃ nanobelts could be controlled by adjusting the size of the Ga droplets.

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