

**Hygroscopicity of
calcium mineral dust
particles**

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Effect of chemical mixing state on the hygroscopicity and cloud nucleation properties of calcium mineral dust particles

R. C. Sullivan^{1,*}, M. J. K. Moore¹, M. D. Petters², S. M. Kreidenweis²,
G. C. Roberts³, and K. A. Prather^{1,3}

¹Dept. of Chemistry and Biochemistry, Univ. of California, San Diego, La Jolla, CA, 92093, USA

²Dept. of Atmospheric Science, Colorado State Univ., Fort Collins, CO, 80523, USA

³Scripps Institution of Oceanography, Univ. of California, San Diego, La Jolla, CA, 92093, USA

* present address: Dept. of Atmospheric Science, Colorado State Univ., Fort Collins, CO, 80523, USA

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Correspondence to: R. C. Sullivan (rsullivan@atmos.colostate.edu)

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Abstract

Atmospheric mineral dust particles can alter cloud properties and thus climate by acting as cloud condensation nuclei (CCN) that form cloud droplets. The CCN activation properties of various calcium mineral dust particles were studied experimentally to investigate the consequences of field observations showing the segregation of sulfate from nitrate and chloride between individual aged Asian dust particles, and the enrichment of oxalic acid in Asian dust. Each mineral's observed apparent hygroscopicity was primarily controlled by its solubility, which determines the degree to which the mineral's intrinsic hygroscopicity can be expressed. The significant increase in hygroscopicity caused by mixing soluble hygroscopic material with insoluble mineral particles is also presented. Insoluble minerals including calcium carbonate, representing fresh unprocessed dust, and calcium sulfate, representing atmospherically processed dust, had similarly small apparent hygroscopicities. Their activation is accurately described by a deliquescence limit following the Kelvin effect and corresponded to an apparent single-hygroscopicity parameter, κ , of ~ 0.001 . Soluble calcium chloride and calcium nitrate, representing atmospherically processed mineral dust particles, were much more hygroscopic, activating similar to ammonium sulfate with $\kappa \sim 0.5$. Calcium oxalate monohydrate ($\kappa = 0.05$) was significantly less CCN-active than oxalic acid ($\kappa = 0.3$), but not as inactive as its low solubility would predict. These results indicate that the common assumption that all mineral dust particles become more hygroscopic and CCN-active after atmospheric processing should be revisited. Calcium sulfate and calcium oxalate are two realistic proxies for aged mineral dust that remain non-hygroscopic. The dust's apparent hygroscopicity will be controlled by its chemical mixing state, which is determined by its mineralogy and the chemical reaction pathways it experiences during transport.

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1 Introduction

Atmospheric aerosols affect the formation, growth, and lifetime of clouds by acting as cloud condensation nuclei (CCN) which provide seeds for cloud droplet formation (McFiggans et al., 2006). Increased number concentrations of CCN can increase cloud brightness (Twomey, 1977), change cloud lifetime (Albrecht, 1989), and suppress (Lohmann and Feichter, 2005; Rosenfeld et al., 2001) or enhance precipitation (Petters et al., 2006; Rosenfeld et al., 2008). The interactions of aerosols and clouds remain the largest uncertainties in our ability to understand and predict the effect that anthropogenic activities have on regional and global climate change (IPCC, 2007). A particle's ability to absorb enough water such that it activates and becomes a cloud droplet is governed primarily by the particle's size and its soluble hygroscopic chemical content (McFiggans et al., 2006). Therefore, the size-resolved chemical composition and mixing state of atmospheric aerosols, and a detailed knowledge of how different compounds interact with water, are required to accurately predict how a realistic aerosol population will undergo cloud nucleation (Andreae and Rosenfeld, 2008; McFiggans et al., 2006).

Most inorganic salts present in aerosols (e.g. $(\text{NH}_4)_2\text{SO}_4$, NaCl) are hygroscopic and CCN-active at sub 100 nm diameters and typical cloud supersaturations. However, the role of insoluble and slightly soluble inorganic salts should also be considered and studied for their CCN activation properties (Kelly et al., 2007). The conversion of insoluble salts to more soluble compounds through heterogeneous and multiphase reactions is of particular interest. Mineral dust particles are the most atmospherically relevant example of these types of inorganic systems. To date, only a few laboratory studies have been reported of the CCN properties of model mineral dust particles (Gibson et al., 2006). Koehler et al. (2007) measured the first CCN activation curves of an authentic dust sample, salt-rich playa dust from Owens (dry) Lake in California. Vlasenko et al. (2005) reported that Arizona test dust had very small subsaturated hygroscopic growth. A few studies have recently been reported probing the role that ambient min-

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eral dust particles play in contributing to aerosol hygroscopicity and CCN formation (Carrico et al., 2003; Perry et al., 2004; Roberts et al., 2006; Shi et al., 2008).

Mineral dust close to the source region is composed of a large and variable mixture of different minerals. This can include insoluble aluminosilicate clays (e.g. kaolinite, montmorillonite, illite), metal oxides (e.g. hematite, silica), carbonates (e.g. calcite, dolomite), and sparingly soluble gypsum ($\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$), as well as soluble halite (NaCl) salts (Jeong, 2008; Krueger et al., 2004; Shi et al., 2005; Sullivan et al., 2007a). These mineral components often undergo chemical reactions during atmospheric transport, such as reactions with $\text{SO}_2(\text{g})$ to produce sulfate, $\text{HCl}(\text{g})$ producing condensed-phase chloride, and nitrogen oxides such as $\text{NO}_2(\text{g})$, $\text{HNO}_3(\text{g})$, and $\text{N}_2\text{O}_5(\text{g})$ to produce nitrate compounds (Krueger et al., 2004; Sullivan et al., 2007a, b; Usher et al., 2003). Reaction with organic compounds such as carboxylic acids by dust is also possible (Falkovich et al., 2004; Sullivan and Prather, 2007). Kelly et al. (2007) recently reported theoretical calculations predicting a wide range of CCN properties due to the differing water soluble fraction of the various components of mineral dust. Mineral dust particles can also accumulate soluble material through internal mixing with sea salt particles during cloud processing (Sullivan et al., 2007b), and condensation of compounds including ammonium nitrate and sulfate.

Calcite ($\text{CaCO}_3(\text{s})$) is generally regarded to be the most chemically reactive mineral component in aeolian dust particles (Laskin et al., 2005; Tang et al., 2004; Usher et al., 2003). It can be found in the form of individual mostly-calcite particles, and also as calcite nanofibers distributed throughout mineral dust particles (Jeong and Chun, 2006). Its strong alkaline nature causes it to react readily with acidic gases including HNO_3 , HCl , and SO_2 . This converts calcium carbonate to other calcium salts such as calcium nitrate, calcium chloride, and calcium sulfate, respectively. The conversion of insoluble calcium carbonate particles to deliquescent calcium nitrate through reaction with nitric acid vapor was first demonstrated in laboratory experiments by Krueger et al. (2004, 2003) and has also been observed in recent field studies (Laskin et al., 2005; Matsuki et al., 2005; Shi et al., 2008). Gibson et al. (2006) demonstrated the large

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effect that this conversion has on the CCN activation of laboratory calcium mineral particles; the critical supersaturation of 100 nm calcium nitrate particles was reduced to 0.11% from ~0.55% for atomized 100 nm calcium carbonate particles.

The addition of soluble material such as nitric and sulfuric acids and their ammonium salts, and the conversion of carbonates to more soluble compounds during atmospheric processing of mineral dust particles is commonly believed to increase the CCN activity of these particles (Fan et al., 2004; Gibson et al., 2006; Levin et al., 1996; Perry et al., 2004; Roberts et al., 2006; Sullivan et al., 2007a). From our single-particle measurements of atmospherically processed mineral dust particles during ACE-Asia, we observed that nitrate and sulfate were not found in relatively large amounts in the same individual dust particles (Sullivan et al., 2007a). We hypothesized that this segregated mixing state of secondary acids could impart very different physicochemical properties to individual aged dust particles, particularly in terms of their interaction with water, due to the large solubility differences between some sulfate and nitrate salts. Oxalic acid was also found to be enriched in the Asian mineral dust particles and we hypothesized that the formation of complexes between oxalate and divalent cations such as Ca^{2+} could reduce the solubility of the dominant fraction of water soluble organic carbon (Sullivan and Prather, 2007). We explore both of these hypotheses herein through laboratory studies of the CCN activation properties of various pure calcium minerals as proxies for fresh and completely processed calcium carbonate particles. The important effect of water and metastable hydrate formation on the hygroscopicity of calcium minerals produced by the particle generation method is the focus of a related manuscript (Sullivan et al., 2009).

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2 Experimental methods

2.1 Aerosol generation

Commercially available calcium salts were used to generate particles for these experiments. They included CaCO_3 (Solvay, uncoated ultrafine Socal[®] 31, unstated purity), CaSO_4 (Aldrich, 99%), $\text{Ca}(\text{NO}_3)_2$ (Alfa Aesar, 99.0+%), $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$ (Aldrich, 98+%), $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$ (Alfa Aesar, 99+%), and $\text{C}_2\text{O}_4\text{H}_2$ (Aldrich, 99+%). The wet-generation method involved atomization of aqueous solutions using high purity N_2 at 2–3 Lpm, and then passing the aerosol through three silica gel diffusion driers to reduce the RH below 5%. Calcium nitrate ($\text{Ca}(\text{NO}_3)_2$), calcium chloride (CaCl_2), and oxalic acid ($\text{C}_2\text{O}_4\text{H}_2$) solutions were made by mixing 0.15 g powder into 150 mL milli-Q water ($>18.2 \text{ M}\Omega$). Dry generation, used for CaSO_4 , CaCO_3 , and $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$, involved placing several grams of powder in an Erlenmeyer flask with a Teflon-coated stir bar. A piece of Teflon tubing was passed through a silicone stopper with a plastic pipette tip fitted on the end. When nitrogen gas was passed through the tubing a jet of gas was produced and directed at the powder, while the stir bar or vortex shaker agitated the dry powder. The suspended aerosol was sampled through a $3/8''$ o.d. stainless steel (s.s.) tube at the top of the flask and then directed into a second Erlenmeyer flask through a s.s. tube with a right angle bend. The remaining aerosol was transferred to a $\sim 17 \text{ L}$ s.s. residence chamber to which dilution nitrogen gas of 2.0–8.0 Lpm was added. The aerosol sample was drawn from two $3/8''$ o.d. s.s. tubes that protruded slightly into the chamber and then combined into one flow.

2.2 CCN activation curve measurement

For both generation methods the dry aerosol was sent to a differential mobility analyzer (DMA, Model 3081, TSI Inc.) for size selection. The aerosol was first passed through a Kr-85 bipolar neutralizer to achieve charge equilibrium. An aerosol sample flow rate of 1.0 or 1.5 Lpm was used, with a sheath flow rate of 10.0 or 15.0 Lpm.

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A 10:1 sheath/sample flow ratio was used whenever possible; however a lower sheath/sample flow ratio was required to select particles as large as 500 nm. The pseudo-monodisperse aerosol was then sent to a flow splitter and transferred to a condensation particle counter (CPC, Model 3010 or 3007, TSI Inc.) and a cloud condensation nuclei counter (CCNc). A small excess aerosol flow of ~ 0.2 Lpm was vented through a HEPA filter to eliminate any large pressure or flow imbalances. Conductive silicone tubing was used to transport the aerosol and minimize wall losses. A scanning mobility particle sizing (SMPS) system (TSI Model 3081 DMA and Model 3010 CPC) was used to verify the mode diameter of the size-selected aerosol.

2.3 CCN counter (CCNc) operation and calibration

The CCNc is a miniature continuous-flow streamwise thermal-gradient CCN chamber constructed at Scripps Institution of Oceanography based on the design of Roberts and Nenes (2005). Briefly, the aerosol sample is introduced into the center of the growth column surrounded by a sheath of filtered humidified air. A linear thermal gradient along the column's continuously wetted wall produces a well-defined and controlled water supersaturation. This supersaturation is primarily a function of the column's thermal gradient (dT), and the total flow rate. Particles that are able to absorb water and grow to optical diameters $>1 \mu\text{m}$ are detected by an optical particle counter (OPC) at the end of the column and are counted as CCN. The aerosol flow rate is 15.0 sccm while the sheath flow is adjusted to produce a total flow of 100.0 sccm. A bypass flow of 400 sccm is used to increase the initial sample flow to the top of the column and reduce particle losses.

Calibration of the CCNc's supersaturation (SS) as a function of column temperature gradient (dT) was performed monthly using dried ammonium sulfate particles atomized from a 1.0 g/L solution of ammonium sulfate (Sigma, 99.999%) and determining activation diameter as a function of column dT . Köhler theory using thermodynamic data from the aerosol inorganic model (AIM) was used to calculate the supersaturation that corresponded to each activation diameter, assuming spherical particles, to create the

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dT–SS calibration curve. Our choice of thermodynamic data for ammonium sulfate corresponds to a single-parameter hygroscopicity of $\kappa=0.61$ (see below). The linear fits of SS to column dT typically had $R^2>0.995$. Column temperature gradients of 2.0 to 15.0°C generated SS of approximately 0.08 to 1.1%.

5 2.4 Data analysis and fitting

CCN activation curves were generated either by scanning across dry particle diameter (D_{dry}) at a fixed SS, or by scanning SS (through changing dT) at a fixed dry diameter. Both the CCNc and CPC record 1-second averaged data, which were averaged to one minute for calculation of the CCN/CN ratio. For both SS and size scans we determined
10 the critical activation diameter (D_{ca}) or critical supersaturation (s_c) by first scaling all ratios to the maximum observed CCN/CN ratio, for large sizes/high supersaturations. The contribution from larger multiply charged particles was then accounted for following Rose et al. (2008). A Boltzman sigmoid curve of the form

$$y = \frac{A_1 - A_2}{1 + e^{(x-x_o)/dx}} - A_2 \quad (1)$$

15 was used to fit each activation curve, where y is the CCN/CN ratio, A_1 is the minimum value, A_2 is the maximum value, dx is the rate of change, and x_o is the midpoint between A_1 and A_2 . The data were fit such that A_1 corresponded to the CCN/CN ratio of the multiply charged particles at small sizes/supersaturations, and A_2 is set to the scaled maximum CCN/CN ratio of ~ 1.0 . Thus, x_o corresponds to the CCN/CN=0.5
20 midpoint typically used to define the critical supersaturation or activation diameter of the singly-charged particles of interest. We note that this technique is more accurate for SS scans, because the fraction of multiply charged particles larger than the selected dry diameter is constant for all points, and the assumption of constant composition with size, implicit in diameter scans, is not as important. SS scans are slower than size
25 scans, however, because the column must equilibrate to the new temperature gradient, and thus size scans were used for some experiments. The contributions of multiply

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charged particles to the activation curve in size scans can be explicitly calculated using the inversion proposed by Petters et al. (2007). Although we did not use that method to reduce all of our size-scan data, we verified via spot checks that our simpler method did not give significantly different results from the full inversion.

5 2.5 Error analysis

The major sources of uncertainty in determining the CCN activation properties, critical supersaturation (s_c) and dry (activation) diameter (D_{dry}), are primarily due to the calibration of the CCNc's supersaturation, and errors in selecting a monodisperse aerosol with the DMA. The SS uncertainty was estimated from the linear fit of the CCNc calibration as 3%. Uncertainty in size selection was estimated as 5%, including variations in particle shape factor and flow rates, and from SMPS measurements of the size-selected aerosol. The error reported for the fitted hygroscopicity parameter, κ (see below), was estimated from one standard deviation of the κ values determined from the individual s_c - D_{dry} pairs for each system.

15 3 Results and discussion

3.1 Hygroscopicity parameter

We have expressed all of our experimental data in terms of the hygroscopicity parameter, κ , introduced by Petters and Kreidenweis (2007), which describes a particle's water activity, a_w , via

$$20 \frac{1}{a_w} = 1 + \kappa \frac{V_s}{V_w}, \quad (2)$$

where V_s is the volume of the dry particle mass and V_w is the volume of water. A choice of dry diameter (D_{dry}) and κ enable calculation of the expected supersaturation – wet

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diameter relationship (“Köhler curve”):

$$S(D) = \frac{D^3 - D_{\text{dry}}^3}{D^3 - D_{\text{dry}}^3(1 - \kappa)} \exp\left(\frac{A}{D}\right), \quad (3)$$

where S is the water saturation ratio over the aqueous droplet, D is the droplet diameter, and $A=2.1 \times 10^{-9}$ m is a constant evaluated for a surface tension of 0.072 J m^{-2} (pure water) and temperature of 298.15 K; the exponential term is the Kelvin term from the Köhler equation. Unless otherwise stated we assume a surface tension of pure water in all cases when deriving κ . The maximum in the curve described by Eq. (3) is identified as the critical saturation ratio for activation of the chosen dry diameter (critical supersaturation, $s_c=S-1$). The best-fit hygroscopicity parameter that describes the observed behavior of test particles can be determined using Eq. (3) by computing Köhler curves for iterations of κ until the modeled s_c for a selected dry diameter matches the experimentally observed s_c .

κ increases with increasing hygroscopicity, from 0 to a maximum of ~ 1.4 ; NaCl has $\kappa=1.28$, for example. We denote the κ value that is derived from measured s_c - D_{dry} pairs and Eq. (3) as apparent hygroscopicity, or κ_{app} , while the κ value that is obtained from Eq. (2) is the intrinsic hygroscopicity, or κ_{intr} , for an aqueous particle. Apparent and intrinsic hygroscopicities are equal when the particle is fully dissolved at the water content relevant to CCN activation. Particles composed of compounds with limited solubility can reduce the observed κ_{app} below κ_{intr} . Similarly, surface tension effects can cause κ_{app} to differ significantly from κ_{intr} .

It is important to properly distinguish between the roles of solubility and hygroscopicity. While hygroscopicity determines the ability of a particle to absorb water, only the compound's dissolved mass fraction contributes to hygroscopic growth. Insoluble components lower the s_c required for activation by increasing the particle's diameter and thus reducing the Kelvin term in Eq. (3), but cannot participate in hygroscopic growth. Therefore, particles composed of intrinsically hygroscopic but insoluble compounds will exhibit a very small apparent hygroscopicity and poor CCN activity. As many of the

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minerals used in this study have limited solubility in water, we applied an extension of Eq. (3) proposed by Petters and Kreidenweis (2008), where the following equations are used to compute κ :

$$\kappa_{\text{sum}}(D) = \sum_i \varepsilon_i \kappa_i H(x_i)$$

$$x_i = \left((D/D_{\text{dry}})^3 - 1 \right) C_i / \varepsilon_i \quad (4a-c)$$

$$H(x_i) = \begin{cases} x_i & x_i < 1 \\ 1 & x_i > 1 \end{cases}$$

Equations (4a–c) have been written to allow for multiple components, i , where ε_i is the volume fraction of component i in the dry particle, C_i the solubility of i in water (expressed as volume compound per unit volume of water), and H is an index for the fraction of component i that has dissolved, x_i . In Eq. (4a) the individual κ_i are intrinsic hygroscopicities for each compound and $\kappa_{\text{sum}}(D)$ is used in Eq. (3) to find the maximum of the Köhler curve. It is important to distinguish $\kappa_{\text{sum}}(D)$ which varies with droplet size, D , and the associated water content, from the constant κ_{app} , which is fit to observed s_c - D_{dry} pairs.

Figure 1 shows representative Köhler growth curves computed from Eqs. (3) and (4) for two of the insoluble minerals studied in this work. The curves demonstrate that, for species with limited solubility, some ranges of D represent water contents that are insufficient to dissolve the entire compound present in the dry particle. Both of these pure 200 nm particles are predicted to activate at $s_c \sim 1\%$, the critical supersaturation required for an insoluble and thus nonhygroscopic but wettable particle ($\kappa_{\text{app}} \sim 0$) of the selected dry diameter. The large s_c is the result of the compound's low solubility which results in the deliquescence relative humidity exceeding the critical supersaturation of an equivalent metastable solution droplet. Such particles thus activate under a deliquescence-limitation (Kreidenweis et al., 2006) and the activation of such insoluble particles is predominantly determined by the Kelvin term in Eq. (3).

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Whether a particle expresses its intrinsic hygroscopicity at activation, or activates at some apparent hygroscopicity depends on κ_{intr} , solubility, particle phase, and initial dry diameter. The apparent hygroscopicity is that which is expected to manifest both experimentally and in the atmosphere. Figure 2 displays the apparent hygroscopicity that would be deduced from a measured s_c - D_{dry} pair as a function of the solubility of the primary compound in the particle. The two cases are for an intrinsic hygroscopicity of $\kappa_{\text{intr}}=1.0$ (left panel) and $\kappa_{\text{intr}}=0.050$ (right panel). The contours show that as the compound's solubility increases a larger fraction of its intrinsic hygroscopicity will be observed as its apparent hygroscopicity, until the solubility limit is exceeded and the full intrinsic hygroscopicity is expressed. We show calculations for three particle sizes. The smallest particle sizes have a reduced apparent hygroscopicity at relatively high solubilities. This is because the higher supersaturations required to deliquesce smaller particles, due to an increased Kelvin effect, are interpreted as lower hygroscopicity (Kreidenweis et al., 2006). Thus, for a solubility at which larger particles may express their $\kappa_{\text{app}}=\kappa_{\text{intr}}$ during activation, the critical supersaturation observed for smaller particles might appear close to the $\kappa_{\text{app}}=0$ isoline (Kelvin limit for insoluble but wettable particles). The larger the intrinsic hygroscopicity of the primary component, the smaller the solubility at which this deliquescence limitation is observed. The presence of sufficiently soluble contaminant that is in solution reduces the solubility where intrinsic and apparent hygroscopicities diverge, as shown by the grey lines in Fig. 2. The presence of the 1% $\kappa_{\text{intr}}=0.6$ contaminant limits the smallest apparent κ value that can be observed experimentally; for the cases considered here, this limit is at $\kappa_{\text{app}}=0.006$ ($=0.6\times 1\%$, Eq. 4a). The contaminant also allows the intrinsic hygroscopicity of the primary component to be expressed at lower solubilities than for single-component particles. This is because the presence of the sufficiently soluble contaminant assures that there is always some water associated with the particle that can dissolve at least a portion of the limited-solubility primary component. We note that the presence of the trace contaminant also modifies the apparent hygroscopicity observed when the particle is fully dissolved at activation. For the case where the primary component

was assumed to have $\kappa_{\text{intr}}=0.05$, the observed hygroscopicity at complete dissolution is slightly larger ($\kappa_{\text{app}}=0.0555$) in the presence of the contaminant, as predicted from the mixing rule for κ (Eq. 4a).

The effect of a 1% by volume $(\text{NH}_4)_2\text{SO}_4$ impurity on the Köhler curve for a 200 nm CaSO_4 particle is shown in Fig. 1. The presence of this soluble and hygroscopic minor component removes the deliquescence-limitation experienced by pure CaSO_4 because the $(\text{NH}_4)_2\text{SO}_4$ fraction deliquesces at subsaturated RH. This ensures that there is always a dissolved component present capable of absorbing water, and also causes some of the CaSO_4 core to dissolve before activation. This reduces s_c from 1.03% to 0.429%, occurring at the first maximum in the growth curve (point I), and κ_{app} at activation is increased from 0 to 0.0075. The second maximum occurs after the CaSO_4 core has fully dissolved (at point II) and corresponds to the critical supersaturation of the completely aqueous $(\text{NH}_4)_2\text{SO}_4/\text{CaSO}_4$ mixed particle (point III). This is analogous to the descriptions of the Köhler curve used to model activation of limited-solubility components in mixed organic-inorganic particles (Bilde and Svenningsson, 2004; Broekhuizen et al., 2004; Petters and Kreidenweis, 2008; Shulman et al., 1996).

Based on Fig. 2 we delineate three solubility regimes: 1) “sufficiently soluble compounds” are those that have $C > \sim 0.1$ and are soluble enough that $\kappa_{\text{app}} = \kappa_{\text{intr}}$ for all dry particle sizes. Therefore CaCl_2 , $\text{Ca}(\text{NO}_3)_2$, and $\text{C}_2\text{O}_4\text{H}_2$ are expected to activate according to κ_{intr} . 2) “Sparingly soluble compounds” are those where κ_{app} can take any value between ~ 0 and κ_{intr} . An approximate range is $10^{-4} < C < 10^{-1}$. 3) “Insoluble compounds” are those for which $\kappa_{\text{app}} \sim 0$ always applies, i.e. $C < \sim 10^{-4}$. CaCO_3 , CaSO_4 , and $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$ (Table 1) are thus not expected to express their large intrinsic hygroscopicities, either experimentally or during cloud nucleation. As particles composed of low solubility compounds undergo reactions/processes both in the laboratory and atmosphere that either convert the compounds into more soluble forms, or add soluble hygroscopic material to the particles, the modified particles will express a larger apparent hygroscopicity.

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The intrinsic hygroscopicity of a limited-solubility compound, expressed once it is fully dissolved, is difficult to measure, but it can be approximated assuming Raoult's Law applies:

$$\kappa_{\text{intr}} = \frac{\nu \cdot \rho_s \cdot M_w}{\rho_w \cdot M_s}, \quad (5)$$

where ρ is the density and M the molecular weight of the solute, s , and of water, w , and ν is the number of ions and molecules produced when one molecule of the solute completely dissociates. Table 1 summarizes the properties of the calcium salts considered here. It includes estimates of the intrinsic and apparent hygroscopicities of each compound, the latter estimated for a 200 nm particle, a typical particle size used in this work, and an amount of soluble contaminant equivalent to the compound's stated purity.

3.2 Experimentally determined hygroscopicities

Representative experimentally determined activation curves for each compound are shown in Figs. 3 and 4, using SS-scans to obtain critical supersaturations for fixed dry diameters. All of the $D_{\text{dry}}-s_c$ data obtained in this study, and the best-fit κ for each data point, are listed in Table 2. Averaged observed values of κ_{app} are summarized for each compound in Table 1.

For calcium chloride, calcium nitrate, and oxalic acid only wet-generated particles were studied (Fig. 3). Because these compounds are sufficiently soluble, we expect $\kappa_{\text{app}} = \kappa_{\text{intr}}$ listed in Table 1. The observed s_c-D_{dry} relationships of these particle types are $\kappa_{\text{app}} = 0.48 \pm 0.07$, 0.51 ± 0.07 , and 0.50 ± 0.05 , respectively, similar to ammonium sulfate ($\kappa_{\text{app}} = \kappa_{\text{intr}} = 0.61$). The particles were likely composed of hydrates (e.g. $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$, $\text{Ca}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$, and $\text{C}_2\text{O}_4\text{H}_2 \cdot 2\text{H}_2\text{O}$) (Kelly and Wexler, 2005) after atomization because the rapid drying process causes those metastable states to be locked in. However, we cannot exclude the possibility that the stable anhydrous stable (Tang and Fung, 1997) or glassy states (Cohen et al., 1987) formed after drying the parti-

cles. Koehler et al. (2006) observed that the hygroscopicity of atomized oxalic acid was generally most consistent with the formation of the dihydrate in typical laboratory atomization/drying processes. As indicated in Table 1, the calculated intrinsic κ values for the hydrated compounds are lower than for the anhydrous compounds, and correspond well to our measurements. Our results also compare reasonably well with those of Gibson et al. (2006) for calcium nitrate. They found $s_c=0.11\%$ and estimated the dry diameter of their 100 nm particles to be 89 nm due to incomplete particle drying, corresponding to $\kappa_{app}=1.25$ (Table 2).

Calcium carbonate, calcium sulfate, and calcium oxalate monohydrate were studied using two different particle generation methods: atomization from an aqueous suspension/solution, and resuspension of the dry powder. The wet-generated particles from these insoluble powders all produced apparent hygroscopicities much larger than the expected κ_{app} considering their very low solubility. The causes of this unexpected behavior are investigated in a related manuscript (Sullivan et al., 2009). For the purposes of this study, only the dry-generated particles that produced the appropriate observed hygroscopicity expected of fresh atmospheric mineral particles are considered here.

Dry-generated CaSO_4 particles exhibited a small $\kappa_{app}=0.0016\pm 0.0005$, and CaCO_3 displayed $\kappa_{app}=0.0011\pm 0.0004$. There was a large fraction of multiply-charged particles in these size-selected aerosols (Fig. 4) due to the dry-generated aerosol's size distribution peaking at a larger diameter than that selected by the DMA. Gibson et al. (2006) measured s_c of atomized 100 nm CaCO_3 to be $\sim 0.55\%$, corresponding to $\kappa_{app}\sim 0.043$ (Table 2). This observed hygroscopicity is much larger than expected for an insoluble mineral (Kelly et al., 2007) ($\kappa_{app}\sim 0$) and was likely caused by the atomization particle generation method they used; we have observed similarly large apparent hygroscopicities ($\kappa_{app}=0.002-0.2$) for atomized CaCO_3 particles (Sullivan et al., 2009). From our results, we conclude that both CaSO_4 and CaCO_3 have a sufficiently low solubility that their CCN activation behavior can be approximated as that of an insoluble component ($\kappa_{app}=0$) which activates according to the Kelvin effect only. Small amounts of sufficiently soluble contaminants (1% by volume) increase apparent hygroscopicity

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above 0, as shown in Figs. 2 and 5. The small κ_{app} of ~ 0.001 for these dry-generated particles indicates that they likely contained only small quantities of sufficiently soluble impurities ($<1\%$), such as inorganic salts.

To investigate the CCN activity of calcium-rich dust particles internally mixed with organic diacids, such as those observed during ACE-Asia (Sullivan and Prather, 2007), activation curves were determined for calcium oxalate monohydrate ($\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$, COH) particles. The dry-generated COH produced broad activation curves (Fig. 4). The reasons for this are unknown, and similar broad activation curves have been observed for other resuspended mineral dust samples (Koehler, 2007). This could be a result of chemical heterogeneity within the dry powder, or large variations in particle shape between particles of the same mobility diameter, causing a range of physical diameters to be selected by the DMA. As a result these curves are more difficult to fit, introducing greater uncertainty into the determination of the critical diameter or supersaturation. Dry-generated COH had an estimated $\kappa_{\text{app}} = 0.048 \pm 0.013$. The small measured hygroscopicity of dry COH compared to oxalic acid ($\kappa_{\text{app}} = 0.50 \pm 0.05$) is attributed to its low solubility (0.0061 g L^{-1}), which is similar to that of calcium carbonate (0.0058 g L^{-1}). Therefore COH should produce particles with hygroscopicity near $\kappa_{\text{app}} \sim 0.001$, similar to that measured for dry calcium carbonate. The particles would have to contain mass fractions of soluble contaminants on the order of $>10\%$ to exhibit $\kappa_{\text{app}} = 0.048 \pm 0.013$. It is possible that dissolved calcium oxalate may be surface active and therefore reduce the surface tension at the air-solute interface, thus decreasing the activation barrier. We calculate that a large reduction in surface tension from 0.072 J m^{-2} (pure water) to 0.029 J m^{-2} would be required to produce particles with $\kappa_{\text{app}} = 0.048$. This reduction in surface tension is on the upper end for surface active organic compounds considered in CCN activation (Facchini et al., 2000). For example, adipic acid, azelaic acid, and succinic acid have surface tensions of 0.064, 0.061, and 0.065 J m^{-2} , while nonanoic acid has a surface tension of 0.030 J m^{-2} (Broekhuizen et al., 2004). However, Nancollas and Wu (1998) measured the surface tension of COH to be 0.0324 J m^{-2} using a thin layer wicking method. As this is close to the surface

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tension we calculated to be necessary to match observations, reduced surface tension may explain the higher-than-expected apparent hygroscopicity of the dry calcium oxalate monohydrate.

Figure 5 compares the CCN properties of the calcium mineral particles described above in s_c - D_{dry} space with κ isolines overlain. Dry-generated CaCO_3 and CaSO_4 demonstrated low apparent hygroscopicity and thus poor CCN activation potential, approaching the $\kappa=0$ isoline. CaCl_2 , $\text{Ca}(\text{NO}_3)_2$, and $\text{C}_2\text{O}_4\text{H}_2$ approach the hygroscopicity of $(\text{NH}_4)_2\text{SO}_4$ ($\kappa_{\text{app}}=\kappa_{\text{intr}}=0.61$). Dry-generated $\text{CaC}_2\text{O}_4\cdot\text{H}_2\text{O}$ has $\kappa_{\text{app}}\sim 0.048\pm 0.013$, its hygroscopicity lies in-between that of the other systems.

In Fig. 5 we also include isolines for pure CaCO_3 and CaSO_4 , and their mixtures with a hygroscopic soluble contaminant with $\kappa_{\text{intr}}=0.6$ and $C=0.43$, representative of ammonium sulfate (and similar to calcium chloride, calcium nitrate, or oxalic acid), that take into account limited solubility (Eq. 4). The volume fraction of contaminant increases from right to left. Since almost no CaCO_3 has dissolved in all these cases, the activation lines obtained for these mixed particles are almost completely determined by the sufficiently soluble substance alone. Our measurements of dry-generated CaCO_3 and CaSO_4 are consistent with the presence of very small volume fractions of soluble hygroscopic material, which produce a detectable increase in the particle's hygroscopicity and a subsequent decrease in its critical supersaturation (Kelly et al., 2007; Roberts et al., 2002). It is typically very difficult to generate uncontaminated insoluble laboratory particles, and this very small degree of $\sim 0.1\%$ contamination that we observed is likely unavoidable and may already be present at the mineral powder's manufacturing source.

Multilayer adsorption models have been recently proposed to describe the activation of insoluble particles such as some of the minerals considered here (Henson, 2007; Sorjamaa and Laaksonen, 2007). Lines of constant hygroscopicity that we derived from the framework of Sorjamaa and Laaksonen (2007) had very shallow slopes in s_c - D_{dry} space that could not be consistently fit to our experimental results for CaCO_3 or CaSO_4 . Alternatively, the formulations suggested by Henson (2007) could be used,

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but the slopes of the best fits using that approach could not be distinguished from the slopes of the s_c - D_{dry} relationships obtained from traditional Köhler theory. The upper bound of $\Delta G=33.6$ kJ/mole given by Henson produces an isoline from his adsorption model corresponding to $\kappa_{\text{app}}\sim 0.001$, which approximates the CaCO_3 or CaSO_4 data.

Therefore, activation of these insoluble minerals can be described simply by the Kelvin effect in traditional Köhler theory, or by Henson's adsorption isotherm model. If activation proceeded via adsorption we do not need to invoke contamination to explain $\kappa_{\text{app}}\sim 0.001$; either mechanism is plausible.

The interplay between chemical mixing state and solubility is summarized in Table 3, which lists the solubilities of various cation-anion pairs for typical mineral dust components. Sparingly soluble and insoluble compounds are indicated with shading. We calculate κ_{intr} for the anhydrous forms in all cases, while recognizing that hydrated states do form for some of the cation-anion pairs shown and that these will have slightly smaller κ_{intr} . The estimated intrinsic hygroscopicities via Raoult's law (Eq. 5) are >0.4 for all these mixing states. However, the large differences in solubility affect whether the observed apparent hygroscopicity of each compound reflects its intrinsic hygroscopicity, or approaches $\kappa_{\text{app}}=0$ during activation. Our experimental observations of calcium mineral particles largely confirm the expected links between solubilities and observed apparent hygroscopicities, for calcium carbonate, sulfate, and oxalate approximately described by $\kappa_{\text{app}}=0$. Although we did not measure them, based on our findings for calcium minerals, we expect the CCN activities of magnesium carbonate and oxalate to also approach $\kappa_{\text{app}}=0$. Other forms of magnesium minerals, calcium nitrate and chloride, and all of the potassium, sodium, and ammonium compounds, are soluble enough that we expect them to exhibit hygroscopic behavior similar to ammonium sulfate, i.e. $\kappa_{\text{app}}=\kappa_{\text{intr}}$. Note that sulfate and oxalate compounds represent atmospherically processed forms of mineral dust and yet are not always soluble enough to express their large intrinsic hygroscopicities. The experimental CCN data summarized in Fig. 5 are in direct agreement with these expected trends in apparent hygroscopicity.

4 Atmospheric implications

The atmospheric processing of mineral dust particles is typically expected to increase their hygroscopicity and CCN activity (Fan et al., 2004; Gibson et al., 2006; Levin et al., 1996; Perry et al., 2004; Roberts et al., 2006; Sullivan et al., 2007a). This is because heterogeneous reactions with reactive gases including nitric, hydrochloric, and sulfuric acids can convert insoluble carbonates into sparingly or sufficiently soluble compounds that play an important role in hygroscopic growth and cloud droplet activation. Condensation or coagulation of soluble secondary aerosol compounds may similarly increase the dust's hygroscopicity. However, field and laboratory studies are required to validate this. From our laboratory measurements of the CCN ability of various calcium minerals, we have demonstrated that one common mineral representative of chemically aged dust particles – calcium sulfate – has poor CCN activation potential, behaving similarly to calcium carbonate which is a proxy for fresh unaged mineral dust. Another proxy for aged dust, calcium oxalate monohydrate, has only moderate CCN activity. Thus, it is possible for mineral dust particles to undergo extensive chemical reactions during atmospheric transport and yet remain non-hygroscopic, due to the formation of low solubility products (Table 3). It is therefore incorrect to assume that all mineral dust particles will become more hygroscopic and CCN-active during the course of atmospheric processing. The predominant chemical reaction and dust mineralogy must be known to reliably predict the resulting hygroscopicity of the aged mineral dust particles. If calcite-rich dust particles instead react with nitrate or chloride precursors, such as NO_x , HNO_3 , N_2O_5 , or HCl , the resulting calcium nitrate and chloride salts are highly hygroscopic and will result in more CCN-active aged dust particles. As mineralogy varies greatly between dust regions, the source of the dust could also have a large influence on its resulting hygroscopicity after atmospheric aging.

Shi et al. (2008) have recently reported the differing hygroscopicities of individual supermicron ambient Asian mineral dust particles collected in Japan. They observed that dust particles that did not contain sulfate or nitrate (i.e. unreacted dust) and sulfate-rich

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dust particles (i.e. aged dust) did not experience hygroscopic growth even at 90% RH. Nitrate-rich dust particles, however, appeared to deliquesce at as low as 15% RH, as has been previously reported (Krueger et al., 2003; Laskin et al., 2005). Our laboratory experiments agree with their field measurements and demonstrate a segregated hygroscopic state caused by the segregated chemical mixing state of sulfate from nitrate and chloride in individual aged dust particles (Sullivan et al., 2007a).

Calcium oxalate monohydrate was observed to be significantly less CCN-active than oxalic acid. Oxalic acid is typically the dominant component of water soluble organic aerosol mass and has been studied as a large potential contributor to aerosol hygroscopicity and CCN activity (Prenni et al., 2001). In the Asian aerosol sampled during ACE-Asia we found that oxalic acid was predominantly internally mixed with mineral dust particles (Sullivan and Prather, 2007). This mixing state could result in the oxalate forming low solubility complexes with divalent cations such as Ca^{2+} that are enriched in Asian dust. This would reduce the solubility of the oxalic acid and this could in turn reduce the apparent hygroscopicity and CCN-activity of oxalate-containing particles, as we observed here. The conversion of calcium carbonate to calcium oxalate through reaction with oxalic acid will also reduce the mass of CaCO_3 available to be converted to more soluble and hygroscopic forms such as CaCl_2 and $\text{Ca}(\text{NO}_3)_2$. A similar effect is produced by the conversion of calcium carbonate to calcium sulfate through reaction with sulfur oxides.

In discussing the CCN activation potential of atmospheric mineral dust particles, we must also consider the role of ammonium internally mixed with dust, which we observed during ACE-Asia (Sullivan et al., 2007a). We found that ammonium was frequently internally mixed with aged mineral dust particles and appeared to have formed through neutralizing secondary acids that had accumulated in the dust. Ammonium was most strongly associated with sulfate, as expected due to the stability of ammonium sulfate. If a small fraction of sulfate is in the form of $(\text{NH}_4)_2\text{SO}_4$ instead of CaSO_4 , this can increase the dust particle's hygroscopicity due to the higher solubility of $(\text{NH}_4)_2\text{SO}_4$ (Fig. 5). During ACE-Asia, sulfate-rich mineral dust particles were primarily associ-

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ated with submicron iron- and aluminosilicate-rich mineral particles, while nitrate and chloride were enriched in supermicron calcite-rich particles (Sullivan et al., 2007a). However, an appreciable fraction of Ca-rich dust particles did contain smaller but still significant sulfate signals. These likely represent particles that are primarily calcium sulfate, while the aged Fe/Al-rich dust particles with larger sulfate and ammonium signals are indicative of dust particles internally mixed with ammonium sulfate. Based on the laboratory CCN data presented above, we predict that these two chemical forms of aged sulfate-dust particles would have very different CCN-activity. The observation of Shi et al. (2008) of aged sulfate-rich mineral dust particles that remained non-hygroscopic is further confirmation that this is a realistic physiochemical state for aged dust particles in the atmosphere. For nitrate- and chloride-rich aged dust particles, however, the precise chemical form of these acid reaction products will be less important because most atmospherically-relevant nitrate and chloride salts are sufficiently soluble and hygroscopic, regardless of their associated cation (Table 3) (Kelly et al., 2007). The addition of trace amounts of water soluble species (e.g. $(\text{NH}_4)_2\text{SO}_4$) exerts the largest increase in CCN activity on dust particles that are initially non-hygroscopic and insoluble. Hence, the chemical pathways associated with atmospheric aging effectively dictate the cloud nucleation potential of atmospheric dust particles. Our findings provide further motivation for the need to determine both the mixing state and relative amounts of various secondary species in individual mineral dust and other particle types to accurately predict the aged aerosol's hygroscopicity and CCN-activity.

In this discussion the important role of particle hygroscopicity in determining other significant particle properties should also be emphasized. Particulate-phase water alters the dust particle's size, refractive index, and shape, which affect the radiative properties of dust and thus their direct climate effects (Satheesh and Moorthy, 2005). The water content of hygroscopic dust particles may provide a medium for multiphase chemical reactions and change the solubility of important components in mineral dust such as iron. Finally, the increased CCN activity of some aged mineral dust particles will also increase their rate of wet scavenging which affects the environmental mass

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balance and transport of mineral dust (Fan et al., 2004). Our single particle measurements during ACE-Asia revealed that mineral dust composed one-quarter of the total sub-500 nm particle counts in remote clean marine environments. This contribution was even larger closer to Asia and was extremely high during an Asian dust storm.

5 The large role that mineral dust particles can play in composing the potential submicron CCN population should therefore be revisited.

5 Conclusions

The CCN activation properties of various pure calcium minerals were measured as proxies for fresh and atmospherically aged calcium-rich mineral dust particles. A single parameter for particle hygroscopicity, κ , was used to interpret the experimental data. A framework for accounting for the limiting effect of solubility on a particle's observed hygroscopicity was introduced through the concepts of apparent and intrinsic hygroscopicity. A particle's apparent hygroscopicity is that which is experimentally observed. It is equal to its intrinsic hygroscopicity if the particle is completely dissolved before activation. Low solubility components limit the full expression of a compound's intrinsic hygroscopicity as its observed apparent hygroscopicity. Changes in surface tension can produce similar differences between apparent and intrinsic hygroscopicity. This framework can be used to describe any complex particle system, including mixtures of inorganic and organic compounds with varying solubility.

20 The particle generation method can cause significant changes in the apparent hygroscopicity of insoluble mineral particles. Atomization of insoluble powders was observed to produce particles with much larger hygroscopicities than expected (Sullivan et al., 2009). Therefore, only the dry-generated insoluble particles and atomized sufficiently soluble particles were considered here to infer the hygroscopicity they will likely display in the atmosphere. Calcium oxalate monohydrate was moderately CCN active and more hygroscopic ($\kappa_{\text{app}}=0.048$) than its solubility and intrinsic hygroscopicity would predict, but still significantly less hygroscopic than oxalic acid ($\kappa_{\text{app}}=0.50$). The

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moderate apparent hygroscopicity could be due to a reduction in surface tension, or the presence of significant amounts of soluble contaminants.

We found two general states of particle hygroscopicity:

1. Sufficiently soluble minerals, i.e. CaCl_2 and $\text{Ca}(\text{NO}_3)_2$, that are hygroscopic ($\kappa_{\text{app}} \sim 0.5$) and have critical supersaturations similar to $(\text{NH}_4)_2\text{SO}_4$ ($\kappa_{\text{app}} = \kappa_{\text{intr}} = 0.61$).
2. Insoluble minerals, i.e. CaCO_3 and CaSO_4 , with low apparent hygroscopicity ($\kappa_{\text{app}} \sim 0.001$), poor CCN activation ability, and critical supersaturations consistent with deliquescence-limited activation (Kelvin limit).

As we have shown herein, the chemical form of the reactive gases that reacted and mixed with dust to produce secondary compounds must be known to accurately predict the hygroscopicity and CCN activity of the resulting mineral mixture. Table 3 demonstrates that it is incorrect to assume that all chemical reactions that process dust in the atmosphere produce soluble reaction products that increase the CCN ability of dust. Insoluble reaction products such as calcium sulfate are frequently observed in aged mineral dust particles, have low CCN-activity, and should also be considered in evaluating the indirect climate effects of dust on cloud properties (Kelly et al., 2007; Shi et al., 2008).

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Table 1. Properties of minerals^a studied in this work, including predicted and measured hygroscopicities.

Compound	ρ g cm ⁻³	M_s g mol ⁻¹	ν	S g L ⁻¹	C vol sol/ vol water	predicted κ		experimental κ_{app}	notes
						intrinsic ^b (Raoult)	apparent ^c		
CaCl ₂	2.15	110.98	3	813	0.38	1.05	1.04	0.48±0.07 (wet) ^d	0
CaCl ₂ ·2H ₂ O	1.85 ^e	147.02	3	813	0.44	0.68	0.68		1
Ca(NO ₃) ₂	2.5	164.09	3	1440	0.576	0.82	0.82	0.51±0.07 (wet)	2
Ca(NO ₃) ₂ ·4H ₂ O	1.86 ^e	236.09	3	1440	0.774	0.43	0.43		0
CaSO ₄	2.96	136.14	2	2.05	6.9×10 ⁻⁴	0.78	0.0084	0.0016±0.0005 (dry)	0
CaSO ₄ ·2H ₂ O	2.32	172.20	2	2.05	8.8×10 ⁻⁴	0.49	0.0078		3
CaCO ₃ (calcite)	2.71	100.09	2	0.0058 ^f	2.1×10 ⁻⁶	0.97	0.0068	0.0011±0.0004 (dry)	4
CaC ₂ O ₄	2.2	128.10	2	0.0061	2.8×10 ⁻⁶	0.62	0.0060	0.048±0.013 (dry)	0
CaC ₂ O ₄ ·H ₂ O	2.12 ^e	146.12	2	0.0061	2.9×10 ⁻⁶	0.52	0.0060		5
C ₂ O ₄ H ₂	1.90	90.04	3	143	7.5×10 ⁻²	1.14	1.14	0.50±0.05 (wet)	3
C ₂ O ₄ H ₂ ·2H ₂ O	1.653	126.07	3	143	8.7×10 ⁻²	0.71	0.71		0

^a Density and solubility taken from the CRC Handbook of Physics and Chemistry, 88th Edition, for 25°C, except where noted below.

^b Calculated using Eq. (5).

^c For 200 nm particles of specified compound, with 0.1%–2% impurity (see notes) having $\kappa_{intr}=0.6$ and infinite solubility.

^d Wet- or dry-generated laboratory particles.

^e Manufacturer stated value.

^f Plummer and Busenberg (1982).

Notes:

0. Assumed purity of 99%; 1. Aldrich, 98+%; 2. Alfa Aesar, 99.0+%; 3. Aldrich, 99+%;

4. Solvay, assumed purity 99%; 5. Alfa Aesar, 99+%

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Table 2. CCN activation properties and hygroscopicities of calcium minerals.

Compound	^a	D_{dry} (nm)	s_c (%)	κ ^b
CaCl ₂ ·2H ₂ O (wet)	S	75.0	0.265	0.46
	S	100.0	0.173	0.45
	S	125.0	0.126	0.44
	D	154.0	0.078	0.58
Ca(NO ₃) ₂ (wet)	S	75.0	0.244	0.55
	S	100.0	0.171	0.47
	S	125.0	0.126	0.44
	D	152.0	0.078	0.64
	^c	89	0.110	1.25
CaSO ₄ (dry)	S	150.0	0.885	0.0024
	S	200.0	0.683	0.0015
	D	213.5	0.640	0.0014
	S	300.0	0.435	0.0013
Solvay CaCO ₃ (dry)	S	150.0	0.938	0.0018
	S	200.0	0.730	0.0011
	D	232.1	0.640	0.0009
	S	250.0	0.580	0.0009
	S	300.0	0.479	0.0008
OMYA CaCO ₃ (wet)	^c	100.0	~0.55 ^c	0.043
CaC ₂ O ₄ ·H ₂ O (dry)	S	100.0	0.487	0.056
	D	123.63	0.363	0.054
	D	140.28	0.363	0.036
	S	150.0	0.253	0.062
	S	200.0	0.224	0.033
C ₂ O ₄ H ₂ (wet)	S	50.0	0.449	0.54
	D	58.6	0.363	0.52
	S	100.0	0.176	0.44

^a Scan type: S, supersaturation scan; D, dry diameter scan.

^b Single-parameter apparent hygroscopicity, κ_{app} , derived from CCN activation curves.

^c Data from Gibson et al. (2006); s_c estimated from their Fig. 3.

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Table 3. Solubilities^a, estimated intrinsic hygroscopicities^b, and measured apparent hygroscopicities of various inorganic salts^c.

	Ca ²⁺	Mg ²⁺	K ⁺	Na ⁺	NH ₄ ⁺
CO ₃ ²⁻	S = 0.0058	1.8 ²⁰	1110	307	1000 ¹⁵
	<i>κ</i> = 0.97 / 0.001 ^d	1.30	0.90	1.30	0.84
SO ₄ ²⁻	2.05	357	120	281	764
	0.78 / 0.002 ^d	0.80	0.83	1.03 / 0.80 ^e	0.72 / 0.61 ^e
NO ₃ ⁻	1440	712	383	912	2130
	0.82 / 0.51 ^d	0.84	0.88	0.96 / 0.88 ^e	0.77 / 0.67 ^e
Cl ⁻	813	560	355	360	395
	1.05 / 0.48 ^d	1.32	0.89	1.34 / 1.28 ^e	1.02
Oxalate ⁻	0.0061 ²⁰	0.38	364 ²⁰	36.1	52.0
	0.62 / 0.048 ^d	0.79	0.42	0.94	0.65

^a All solubilities (g L⁻¹) taken from the CRC Handbook of Chemistry and Physics, 88th edition, for 25°C. Superscript indicates solubility at *T* other than 25°C.

^b Calculated with Eq. (5) using values listed in Table 1 or else in the CRC Handbook for the anhydrous state.

^c Laboratory generated particles may be composed of salts in metastable states in some cases.

^d Measured apparent hygroscopicity from this work (in italics).

^e Hygroscopicity measured from CCN activation experiments summarized by Petters and Kreidenweis (2007).

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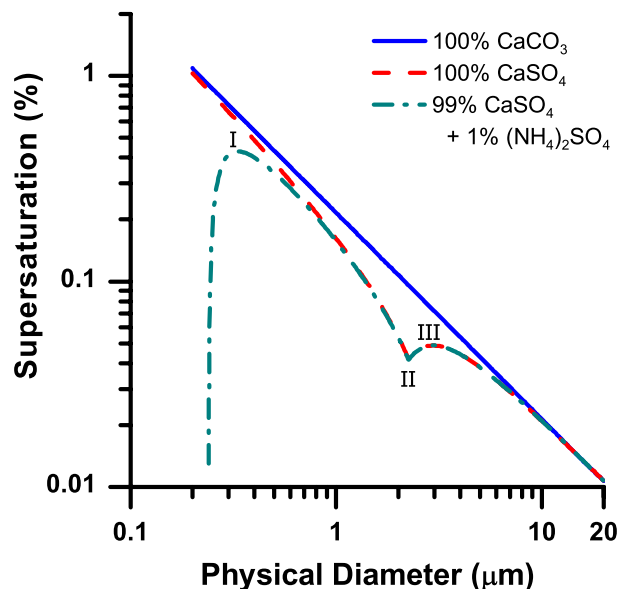


Fig. 1. Theoretical κ -Köhler droplet growth curves for 200 nm dry CaCO_3 and CaSO_4 particles calculated using Eqs. (3) and (4). A mixed 200 nm particle composed of 99% CaSO_4 and 1% $(\text{NH}_4)_2\text{SO}_4$ ($\kappa_{\text{intr}}=0.6$, $C=0.43$) by volume is also shown. The solubilities and intrinsic hygroscopicities (κ_{intr}) used are from Table 1. I: activation of the mixed particle before complete dissolution of the CaSO_4 core, II: complete dissolution of CaSO_4 , III: second supersaturation maximum corresponds to s_c of an equivalent metastable liquid droplet. s_c of pure CaCO_3 and $\text{CaSO}_4 = 1.03\%$, while the mixed CaSO_4 particle has $s_c=0.429\%$.

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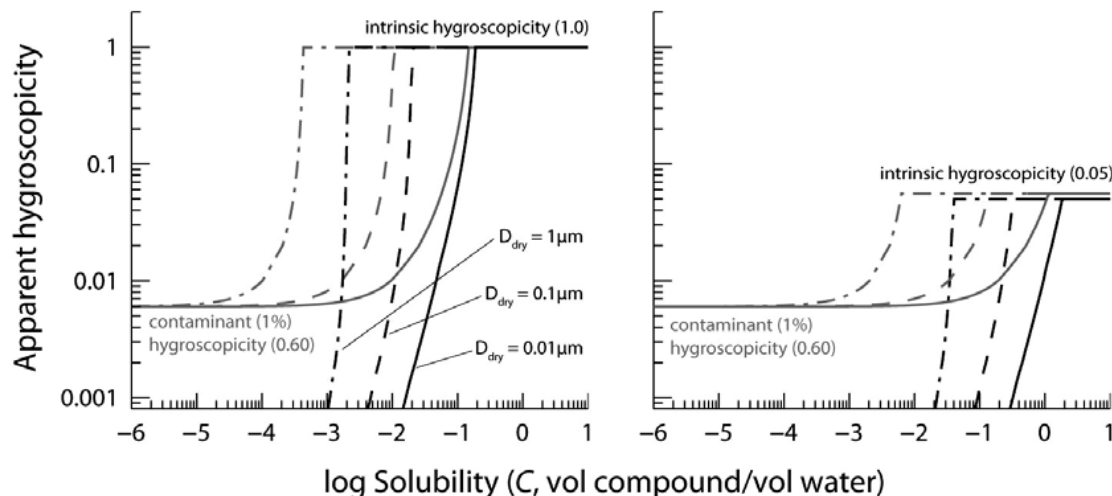


Fig. 2. Predicted apparent hygroscopicity contours for particles composed of a principle compound with $\kappa_{\text{intr}}=1.0$ (left) or 0.050 (right) as a function of the compound's solubility, C , and dry particle size. Grey contours are for mixed particles composed of 99% volume principle compound and 1% contaminant with $\kappa_{\text{intr}}=0.60$ and $C=\infty$.

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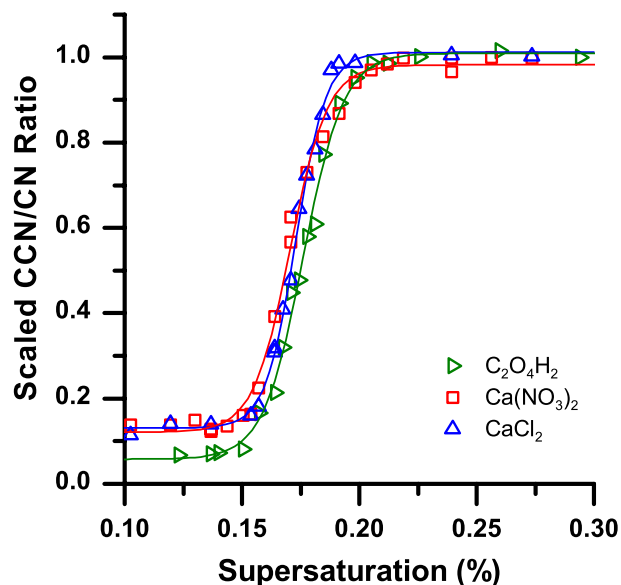


Fig. 3. CCN activation curves for atomized oxalic acid ($s_c=0.176\%$), calcium chloride ($s_c=0.173\%$), and calcium nitrate ($s_c=0.171\%$) from supersaturation scans of 100.0 nm mobility diameter particles. Lines are sigmoidal fits between the scaled lower and upper plateaus. The CCN column's supersaturation is derived from the column's temperature gradient by calibration with ammonium sulfate aerosol.

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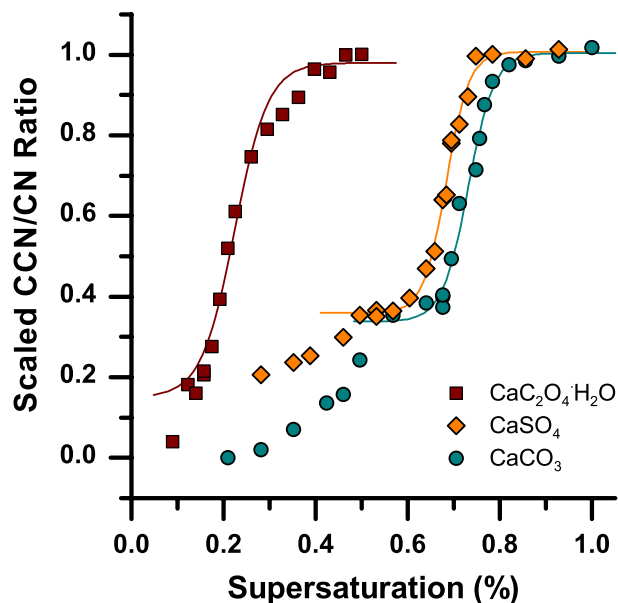


Fig. 4. CCN activation curves of dry-generated calcium carbonate, calcium sulfate, and calcium oxalate monohydrate aerosol obtained by scanning supersaturation of 200.0 nm dry mobility diameter particles. Lines are sigmodal fits to each curve. Note the lack of a clear multiply-charged plateau in the calcium oxalate monohydrate curve.

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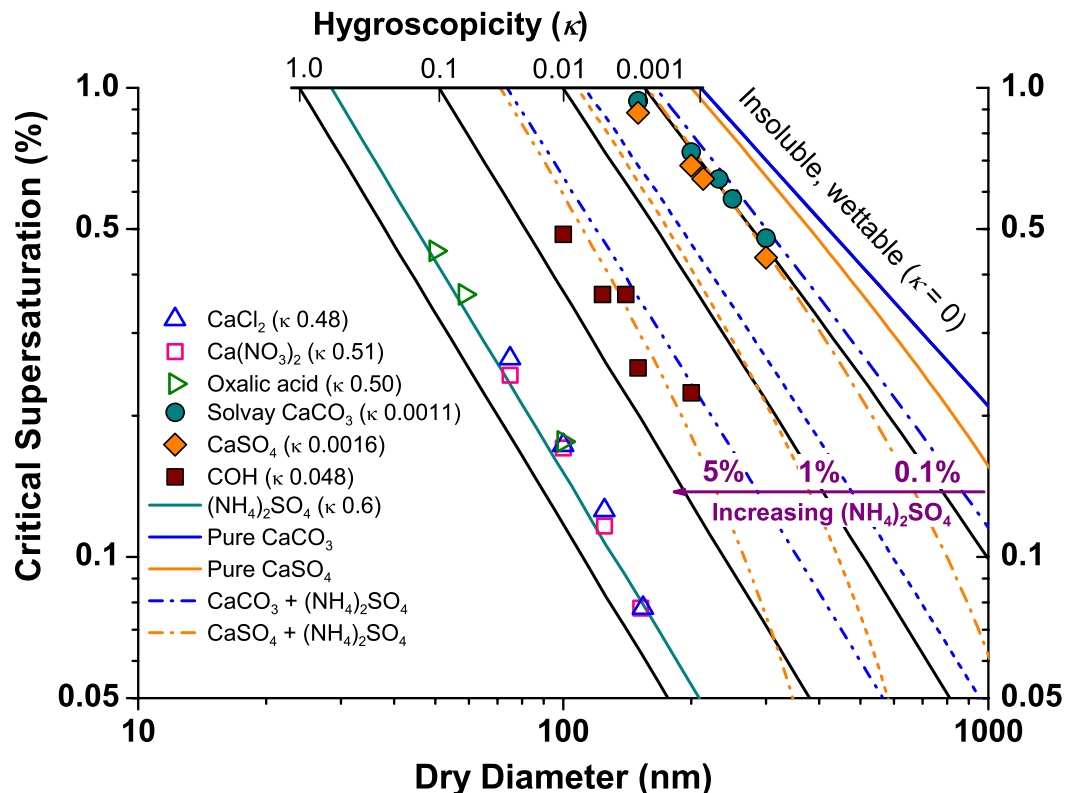


Fig. 5. s_c - D_{dry} relationships with κ isolines overlain for dry-generated insoluble calcium minerals (solid symbols) and atomized soluble salts (open symbols). The best-fit κ_{app} values from Table 1 are also listed. Also displayed are predicted κ isolines for pure CaCO_3 particles (blue lines, $\kappa_{\text{intr}}=0.97$, $C=2.1 \times 10^{-6}$) and CaSO_4 particles (orange lines, $\kappa_{\text{intr}}=0.78$, $C=6.9 \times 10^{-4}$), and their internal mixtures (dashed lines) with varying volume fractions of ammonium sulfate ($\kappa_{\text{intr}}=0.6$, $C=0.43$). 100% CaCO_3 lies along the $\kappa=0$ isoline, and the fraction of $(\text{NH}_4)_2\text{SO}_4$ increases from right to left. The theoretical isoline for pure $(\text{NH}_4)_2\text{SO}_4$ ($\kappa_{\text{app}}=0.6$) is also shown.

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