

# Ternary solution of sodium chloride, succinic acid and water – surface tension and its influence on cloud droplet activation

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Received: 10 March 2008 – Accepted: 14 March 2008 – Published: 14 April 2008

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Published by Copernicus Publications on behalf of the European Geosciences Union.

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## Abstract

Surface tension of ternary solution of sodium chloride, succinic acid and water was measured as a function of both composition and temperature by using the capillary rise technique. Both sodium chloride and succinic acid are found in atmospheric aerosols, the former being main constituent of marine aerosol. Succinic acid was found to decrease the surface tension of water already at very low concentrations. Sodium chloride increased the surface tension linearly as a function of the concentration. Surface tensions of both binary solutions agreed well with the previous measurements. Succinic acid was found to lower the surface tension even if sodium chloride is present, indicating that succinic acid as a surface active compound tends to concentrate to the surface. An equation based on thermodynamical relations was fitted to the data. As a result, a surface tension parameterization of ternary solution was obtained over the whole concentration range. The parameterization can safely be used at temperatures from 10 to 30°C. These kinds of parameterizations are important for example in atmospheric nucleation models. To investigate the influence of surface tension on cloud droplet activation, the surface tension parameterization was included in an adiabatic air parcel model. Usually in cloud models the surface tension of pure water is used. Simulations were done for characteristic marine aerosol size distributions consisting of the considered ternary mixture. We found that by using the surface tension of pure water, the amount of activated particles is underestimated up to 8% if particles contain succinic acid and overestimated it up to 8% if particles contain only sodium chloride. The surface tension effect was found to increase with increasing updraft velocity.

## 1 Introduction

Indirect effect of aerosol particles to the atmosphere (the cloud albedo effect) is currently the most uncertain piece in the climate change puzzle (IPCC, 2007). In order to understand quantitatively how aerosols and clouds interact, the physico-chemical

ACPD

8, 7189–7216, 2008

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properties of aerosols have to be known. As the atmospheric aerosols have highly heterogeneous properties and there are only limited set of measurements characterizing them, these properties have to be treated in approximative manner in climate models. For example in cloud models surface tension of pure water is often used (Anttila et al., 2002).

It is known that a large fraction of aerosol particles is inorganic salts (Dusek et al., 2006; McFiggans et al., 2005). Owing to their hygroscopicity, salt particles act as efficient cloud condensation nuclei (CCN). Measurements of chemical composition of aerosol particles have shown also that, in addition of inorganic salts, an extensive number of both water-soluble and water-insoluble organic acids are present in the aerosol phase (Legrand et al., 2007). They are formed for example in gas to particle phase transitions by oxidation of organic gases (Anttila et al., 2004; Kulmala et al., 2004; Tunved et al., 2007). Soluble organic acids have also found to be able to act as moderately active CCN (Cruz et al., 1997; Bilde et al., 2004; Varga et al., 2007).

Surface active compounds are compounds that tend to concentrate to the surface of a liquid solution and therefore lower the surface tension of the solution. Organic acids are often surface active. According to Köhler theory (Köhler, 1936), a decrease in surface tension induces a decrease in the critical supersaturation of a droplet. This means that organic acids can enhance the cloud droplet activation of mixed particles by decreasing the surface tension of the droplet (Cruz and Pandis, 1997, 1998; Li et al., 1998). The effect of various organic acids on the surface tension of water has been studied earlier (Shulman et al., 1996; Tuckermann and Cammenga, 2004; Hyvärinen et al., 2006). There are also few measurements on the surface tensions of solutions containing both inorganic salt and organic acid (Bilde et al., 2004; Topping et al., 2006; Tuckermann, 2007), but the data is scarce and cover usually only limited temperature and/or concentration range. Previous surface tension measurements show that the presence of inorganic salts can in some cases even enhance the surface activity of organic acids (Tuckermann, 2007; Kiss et al., 2005). The CCN number concentration has also been found to be quite sensitive to the decrease of surface tension caused by

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the presence of organic acids (Chang et al., 2007; Roberts et al., 2002).

There are only a few extensive data sets on surface tensions of ternary mixtures of atmospheric relevance. To this end, we studied the surface tension of ternary solution of sodium chloride, succinic acid and water as a function of both concentration and temperature. Succinic acid is a typical dicarboxylic acid found in aerosol phase (Legrand et al., 2007), and sodium chloride is a hygroscopic compound found in aerosols of marine origin (Heintzenberg et al., 2000). This ternary solution is a good model mixture when investigating atmospheric processes. The measured surface tensions were used as an input to an adiabatic air parcel model (Anttila et al., 2002) to study the effect of the surface tension on cloud droplet activation.

## 2 Materials and methods

### 2.1 Experimental setup

The surface tension was measured using a capillary rise technique (Fig. 1). The sample was placed in a small glass (about 20 ml) in the bottom of a double-walled glass. The temperature of the glass was controlled by circulating thermostatically controlled liquid between the walls (Lauda RC6 CS). To ensure effective heat transfer, the sample vessel was surrounded with water.

The samples were made of pure substances with purity of 99% or better (Table 1). Mixture was made in a 50 ml glass bottle and the weight of each compound was measured using a weighing balance Precisa junior 310M with uncertainty of about 0.1% (absolute accuracy  $\pm 0.003$  g). This corresponds to an uncertainty of less than 0.5% in mole fraction. Both sodium chloride and succinic acid were dried before weighing in an oven at temperature of about 100°C to evaporate all the volatile impurities. Before every measurement the capillary tubes were cleaned using 96% sulfuric acid and purified water (Milli-Q).

The height of the liquid in the capillary tube was measured using a slide caliper

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modified to this purpose. Accuracy of the caliper is  $\pm 0.02$  mm. A pointer made of thin plates of platinum and steel was mounted to the caliper to measure the distance from the surface of the bulk liquid to the bottom of the meniscus, normal to the capillary wall. Only the platinum plate was in contact with the sample. The measurements were performed inside temperature range 283.15 K to 303.15 K. The meniscus was allowed to both rise and descent by altering the pressure in the tube. It was always checked if the meniscus stabilized in the same place when it was let to rise or descent. Uncertainty of the surface tension measurement was estimated to be less than 1%.

## 2.2 Calibration

Two similar glass capillary tubes were used in the measurement. Both of them were calibrated using purified water (Milli-Q). The height of the water inside the tube was measured as a function of temperature. The surface tension can be calculated using equation (Bikerman, 1947):

$$\sigma = \frac{1}{2} \rho g r \left( h_0 + \frac{r}{3} \right), \quad (1)$$

where  $\sigma$  is surface tension,  $\rho$  is density,  $g$  is acceleration due to gravity,  $h_0$  is the height of the water surface inside the capillary tube and  $r$  is the radius of the tube. Because the surface tension and the density of pure water are well known (see Table 1), the radius of the tube can be calculated using Eq. (1). Water was used for calibration, because all the measured solutions were aqueous having surface tensions relatively close to the surface tension of pure water, thus resulting in similar capillary rising heights.

Radii of the capillary tubes are  $198.37 \mu\text{m}$  and  $179.62 \mu\text{m}$ , with standard deviations of 0.90 and 0.62, respectively. Standard deviations were used as errors of the capillary tubes radiuses when calculating the error of the measured surface tension.

### 2.3 Density of the ternary solution

When measuring surface tension using the capillary rise technique, the density of the solution has to be known. We were not able to find literature data of densities of our ternary mixture. Therefore the density had to be estimated using known binary solution densities. The density of aqueous sodium chloride ( $\rho_{13}$ ) as a function of concentration and temperature has been reported earlier by Novotny and Söhnel (1988) (International Critical Tables, 1928; Table 1).

According to previous measurements on density of aqueous succinic acid (Hyvärinen et al., 2006), the density can be estimated by using the definition for the density of ideal solution (Prausnitz et al., 1986):

$$\frac{1}{\rho_{12}} = \frac{M_1 x_1}{\rho_1 [M_1 x_1 + M_2 (1 - x_1)]} + \frac{M_2 (1 - x_1)}{\rho_2 [M_1 x_1 + M_2 (1 - x_1)]} \quad (2)$$

where  $\rho$  is density,  $M$  is the molar mass and  $x$  is the mole fraction. Indices 1 and 2 stand for water and succinic acid, respectively. The assumption of ideality is reasonable due to the fact that succinic acid does not dissociate and form ions when dissolved in water, and due to the low concentration of succinic acid (Bilde and Svenningsson, 2004; Varga et al., 2007). Density of pure succinic acid was estimated using the Yen-Woods method (Yen and Woods, 1966) by Hyvärinen et al. (2006) (Table 1).

Density of the ternary solution can be estimated by taking a weighted average of the binary densities:

$$\rho_{123} = \frac{w_2 \rho_{12}(x_t) + w_3 \rho_{13}(x_t)}{w_2 + w_3}, \quad \text{where } (x_t = x_2 + x_3), \quad (3)$$

$w$  is mass fraction and indices 1, 2, and 3 stand for water, succinic acid and sodium chloride, respectively. This kind of estimation can be considered accurate within the solubility limits, because the effect of succinic acid on the density of the mixture is small compared to the effect of sodium chloride, and the density of aqueous sodium chloride is well known. The error of the ternary density was estimated to be less than 1%.

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## 2.4 Cloud model

The cloud model used in this study is a box model that simulates interactions between an aerosol population and water vapour inside an adiabatically rising air parcel (Anttila et al., 2002). A sectional representation is used to model the aerosol population, which is assumed to be internally mixed, inside the air parcel. A detailed description of the model can be found in reference Anttila et al. (2002).

## 3 Results

### 3.1 Surface tension

Surface tension of the ternary solution of sodium chloride, succinic acid and water was measured as a function of mixture composition and temperature. The range of concentrations covered by measurements was limited by their solubilities. The maximum mole fractions used were 0.0075 and 0.06 for succinic acid and sodium chloride, respectively. The measured surface tensions are presented in Table 2.

#### 3.1.1 Surface tensions of the binary solutions

In Fig. 2 the surface tension measurements of aqueous succinic acid is presented and compared to the previous measurements made with the Wilhelmy-plate method by Hyvärinen et al. (2006). It can be seen that succinic acid decreases the surface tension of the solution already at quite small mole fractions. At succinic acid mole fraction of 0.005 the surface tension is lowered by 3.0 mN/m from the surface tension of water. Results agree well with the previous measurements done using a different method. Surface tensions measured with capillary rise technique are systematically slightly higher, but the difference is not significant. Surface tension of aqueous sodium chloride is presented in Fig. 3. Sodium chloride increases surface tension linearly as

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a function of concentration. Results are compared to literature values (International critical tables, 1928), and they agree well.

### 3.1.2 Surface tension of the ternary solution

In Fig. 4 the surface tension of the ternary solution is presented as a function of mole fractions of succinic acid and sodium chloride at 25°C. The decrease of the surface tension as a function of succinic acid mole fraction is steeper for solutions with larger amount of the salt. This indicates that succinic acid, as a surface active compound, tends to fill the surface of the ternary solution. It has been found in previous studies that in some cases inorganic salts can even enhance the surface tension lowering of organic acids (Tuckermann, 2007; Kiss et al., 2005).

In Fig. 5 the surface tension of the ternary solution is presented as a function of temperature for four different mole fractions of succinic acid. The sodium chloride mole fraction is 0.02. Surface tension of the solution decreased with increasing temperature as expected. The behaviour is similar also with other concentrations of sodium chloride. The gradient of surface tension with respect to the temperature gets lower with increasing concentrations of both solvents. This means that the temperature dependency of the surface tension is a function of the composition of the solution.

### 3.1.3 Surface tension parameterization

To make the measurement results useful for numerical models, they need to be extrapolated to a larger concentration and temperature range. To this end, a function by Chunxi et al. (2000) was fitted to the data. The equation is based on thermodynamical relations and can be written as (Chunxi et al., 2000):

$$\sigma = \sum_i x_i \sigma_i - RT \sum_i \frac{x_i}{\sum_j x_j \Lambda_{ij}} \sum_j x_j \left( \frac{\partial \Lambda_{ij}}{\partial A} \right)_{T,P,x} \quad (4)$$

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where

$$\Lambda_{ij} = \exp\left(-\frac{U_{ij} - U_{ii}}{RT}\right) \quad (5)$$

and

$$\left(\frac{\partial \Lambda_{ij}}{\partial A}\right)_{T,P,x} = -\frac{\Lambda_{ij}}{RT} \left(\frac{\partial (U_{ij} - U_{ii})}{\partial A}\right)_{T,P,x} \quad (6)$$

5 In Eqs. (4)–(6),  $\sigma_i$  is the surface tension and  $x_i$  is the mole fraction of compound  $i$ ,  $U_{ij} - U_{ii}$  is the difference in interaction energies of molecule pairs  $ij$  and  $ii$ ,  $T$  is the temperature,  $A$  is the surface area and  $R$  is the universal gas constant. Parameters

$$U_{ij} - U_{ii} \text{ and } \left(\frac{\partial (U_{ij} - U_{ii})}{\partial A}\right)_{T,P,x} \quad (7)$$

were used to fit the equation to the experimental data.

10 Few assumptions had to be made in order to obtain a fit that represents surface tension of the ternary solution beyond the solubility limits (Table 1). The surface tension of pure succinic acid was estimated using Magleod-Sugden method (e.g. Reid et al., 1987) as applied by Gaman et al. (2004). This method estimates surface tensions of hydrogen bonded liquids with errors typically less than 10%. A linear equation was  
15 fitted to the temperature dependent data (Table 1). This surface tension can be interpreted as the surface tension of succinic acid as subcooled liquid. Sensitivity analysis by Gaman et al. (2004) shows that the accuracy of the method is reasonable for atmospheric nucleation calculations. In absence of relevant data, the surface tension for pure sodium chloride was estimated from molten salt surface tensions (Janz et al.,  
20 1980), and the results were extrapolated to the temperature range considered in this study. In Fig. 6, the behaviour of the surface tension parameterization is presented for the whole concentration range at 25°C. Parameters for the Eq. (1) are presented in Table 3.

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The fit is presented in Figs. 2–6 as solid lines. The standard deviation of the residuals is 0.23 mN/m. The fit is based on measurements conducted at temperatures from 10 to 30°C, but it can be extrapolated with caution to somewhat higher and lower temperatures. The fit presented here can also be used beyond solubility limits with few cautions. No macroscopic phenomena, such as critical micelle concentration, are taken into account. It has to be also noticed that the surface tensions of pure compounds are estimates, and consequently error are prone to increase with increasing concentration of a solvent. Despite these uncertainties, the extended parameterization is applicable in models simulating microscopic phenomena such as nucleation.

## 3.2 Cloud droplet activation

### 3.2.1 Critical supersaturation of sodium chloride/succinic acid particles

Saturation ratio of water vapour over solution droplet can be calculated using Köhler theory (Köhler, 1936):

$$S = a_w \exp\left(\frac{2\sigma M_w}{r\rho_w RT}\right), \quad (8)$$

where  $a_w$  is the water activity,  $\sigma$  is the surface tension of the solution,  $M_w$  is the molecular mass of water,  $r$  is the radius of the droplet,  $\rho_w$  is the density of water,  $R$  is the universal gas constant and  $T$  is the temperature. In Eq. (5) the exponential term accounts for the curvature effect (Kelvin effect) and the activity for the solute effect (Raoult effect). In this study we concentrate on the effect of surface tension and assume that activity  $a_w$  equals to the mole fraction of water  $x_w$ .

The maximum value of  $S$  is called the critical supersaturation ( $S_{\text{crit}}$ ). If the saturation ratio of the environment exceeds the critical value, the droplet can activate and grow to larger sizes. Because the Köhler theory accounts for the surface tension of the solution, the difference of the critical supersaturations due to the surface tension used in the equation was investigated. The relative difference of critical supersaturations

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$[S_{\text{crit}}(\sigma_w)/S_{\text{crit}}(\sigma_{\text{exp}})-1]$  is presented in Fig. 6 as a function of the dry particle diameter. The critical supersaturation is higher if the surface tension of pure water ( $\sigma_w$ ) is used, for all droplets except those containing only sodium chloride. Because surface tension is by definition the energy needed to increase the surface area, the critical supersaturation decreases with decreasing surface tension. It can be seen that surface tension has an effect only when the initial particles are smaller than 100 nm in diameter. The difference increases with increasing mass fraction of succinic acid.

### 3.3 Cloud model simulations

To see if surface tension has an effect on cloud droplet activation, we compared results from cloud model simulations that used two different surface tensions in  $S$ : that of pure water and that predicted by the developed surface tension parameterization. The parameter values describing aerosol size distribution used in the simulations are shown in Table 4. The distribution represents a typical aerosol size distribution measured in marine environment (Heintzenberg et al., 2000). Succinic acid is treated as a completely soluble compound despite of its solubility limit (88 g/l; Saxena et al., 1996). Previous studies have shown that the solubility of a slightly soluble compound has an effect only when there is no or very little of sodium chloride in the particle (Bilde et al., 2004). Also, particles have usually undergone different kinds of humidity conditions in the atmosphere. This means that the solution can be supersaturated with respect to succinic acid, and the assumption of complete solubility is sensible.

Cloud model simulations were made using three different updraft velocities; 0.1, 0.5 and 1.0 m/s. In Fig. 8 the ratio of the number of cloud droplets to the total number of particles (activated fraction) is presented as a function of the initial mass fraction of succinic acid in the aerosols. The activated fraction decreases as the mass fraction of succinic acid increases. This is because succinic acid has higher molar volume than sodium chloride and because sodium chloride is able to dissociate in water.

In Fig. 8, it can be seen that there is a difference in the number of cloud droplets arising from the choice of the surface tension. The difference increases with increas-

ing mass fraction of succinic acid and with increasing updraft velocity. By using the surface tension of pure water in the cloud model the amount of activated particles is underestimated up to 8% for aerosol size distributions containing succinic acid and overestimated up to 8% for size distributions containing only sodium chloride. This is because succinic acid decreases the surface tension of the solution while sodium chloride increases it. When the updraft velocity of the air parcel is 1.0 m/s, the effect of surface tension can clearly be seen with particles having mass fraction of succinic acid over 40%. Mass fractions of organic compounds in atmospheric aerosol particles can easily exceed this value (Dusek et al., 2006). The effect of updraft velocity can be explained by a larger dry diameter of activated particles, which is due to the lower updraft velocity, and thereby lower supersaturation. This decreases the Kelvin effect, which takes into account the surface tension of the solution. This was also observed for marine aerosol size distribution previously by Nenes et al. (2002).

## 4 Conclusions

In this study the surface tension of ternary solution of sodium chloride, succinic acid and water was measured using the capillary rise technique. Both of these substances are found in atmospheric aerosols. Measured surface tensions of binary solutions agreed well with literature values, confirming that the method is applicable to these solutions. Measurements were performed within the concentration range defined by the solubility limits of the solvents. To estimate the surface tension beyond these limits, an equation developed by Chunxi et al. (2000) was fitted to the data. As a result a parameterization of surface tension of the ternary solution was obtained over the whole concentration range. The parameterization is applicable inside temperature range of 10 to 30°C, but it can be extrapolated beyond these limits. There are many benefits with these kinds of parameterizations. First, it gives the surface tension of ternary system as a function of both composition and temperature. Second, the parameterization is based on thermodynamics having fit parameters which can be, for example modelled. This

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gives many options for future research. Because the parameterization can be used beyond solubility limits, it can be also applied, for example, in numerical calculations regarding nucleation or cloud droplet formation.

To estimate the atmospheric relevance of the surface tension data, cloud droplet activation simulations were performed with particles of different composition. Three different updraft velocities (0.1, 0.5 and 1.0 m/s) were used. It was found that while sodium chloride particles act as efficient cloud condensation nuclei, succinic acid can enhance the activation further by decreasing the surface tension of the aqueous solution. By using the surface tension of pure water in the cloud model the amount of activated particles was underestimated up to 8% if the initial particles contained succinic acid. For pure sodium chloride particles it was overestimated up to 8%. Although the changes are small percent-wise, they still may have significance to the cloud radiative properties, in particular over the oceans. The marine clouds are usually clean clouds having lower number concentration of cloud droplets, which makes them especially sensitive to changes in the cloud droplet concentration (Platnic and Twomey, 1994).

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**Table 1.** Properties of the compounds used in this study.

Compound	Chemical composition	<i>M</i> (g/mol)	Solubility in water at 20°C (g/l)	Density of aqueous solution (mN/m)	Surface tension of pure compound (mN/m)	Purity	Manufacturer
Sodium Chloride	NaCl	58.44	358.9 <sup>d</sup>	$\rho_w + 44.85c - 0.096ct + 0.614 \cdot 10^{-3} \cdot ct^2 - 2.712c^{3/2} + 0.01c^{3/2}t^e$	$191.16 - 71.88 \cdot 10^{-3} T^c$	99.5%	Carlo Erba Reagenti
Succinic acid	C <sub>4</sub> H <sub>6</sub> O <sub>4</sub>	118.09	~ 88 <sup>a</sup>	see Eq. (2)	$83.45 - 0.127^b$	ACS quality	J. T. Baker
Water	H <sub>2</sub> O	18.015	–	$999.65 + 0.204t - 6.174 \cdot 10^{-2} t^{3/2} e$	$93.66 + 9.13 \cdot 10^{-3} T - 0.28 \cdot 10^{-3} T^2$	Milli-Q	Milli-Q

*(t)* = °C;*(T)* = K;*(c)* = mol/l.<sup>a</sup> Saxena et al., 1996.<sup>b</sup> Hyvärinen et al., 2006.<sup>c</sup> Janz et al., 1980.<sup>d</sup> CRC, 1998.<sup>e</sup> Novotny and Söhnel, 1988.

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**Table 2.** Surface tensions ( $\sigma$ ) of ternary solution of sodium chloride, succinic acid and water.  $T$  is temperature,  $x$  is mole fraction and subscripts 2 and 3 represent succinic acid and sodium chloride, respectively.

$t$ (°C)	$x_3$	$x_2$	$\sigma$ (mN/m)	$t$ (°C)	$x_3$	$x_2$	$\sigma$ (mN/m)	$t$ (°C)	$x_3$	$x_2$	$\sigma$ (mN/m)	$t$ (°C)	$x_3$	$x_2$	$\sigma$ (mN/m)	$t$ (°C)	$x_3$	$x_2$	$\sigma$ (mN/m)
10.2	0	0	73.5	15.2	0	0	73	20.2	0	0	72.3	25.1	0	0	71.6	30.1	0	0	71
10.3	0	0.0015	72.7	15.3	0	0.0015	72.1	20.2	0	0.0015	71.5	25.2	0	0.0015	70.7	30.1	0	0.0015	70.1
10.4	0	0.0015	72.4	15.3	0	0.0015	71.9	20.2	0	0.0015	71.3	25.2	0	0.0015	70.7	30.2	0	0.0015	70.1
10.3	0	0.003	71.8	15.3	0	0.003	71.2	20.2	0	0.003	70.7	25.2	0	0.003	70	30.2	0	0.003	69.4
10.3	0	0.0045	70.8	15.2	0	0.0045	70.4	20.2	0	0.0045	69.8	25.2	0	0.0045	69.2	30.2	0	0.0045	68.7
10.4	0	0.006	70.1	15.2	0	0.006	69.6	20.2	0	0.006	69.2	25.2	0	0.006	68.7	30.1	0	0.006	68.1
10.3	0	0.0075	69.3	15.3	0	0.0075	68.9	20.2	0	0.0075	68.4	25.1	0	0.0075	67.9	30.2	0	0.0075	67.5
10.2	0.01	0	75.4	15.3	0.01	0	74.5	20.2	0.01	0	73.9	25.1	0.01	0	73.3	30.1	0.01	0	72.7
10.2	0.02	0	75.7	15.2	0.02	0	75.2	20.2	0.02	0	74.5	25.1	0.02	0	73.9	30.1	0.02	0	73.3
10.2	0.02	0.0015	73.9	15.3	0.02	0.0015	73.3	20.3	0.02	0.0015	72.8	25.1	0.02	0.0015	72.3	30.1	0.02	0.0015	71.7
10.3	0.02	0.003	72.6	15.2	0.02	0.003	72.2	20.2	0.02	0.003	71.8	25.1	0.02	0.003	71.3	30.2	0.02	0.003	70.8
10.4	0.02	0.0043	71.7	15.3	0.02	0.0043	71.3	20.2	0.02	0.0043	70.9	25.1	0.02	0.0043	70.4	30.1	0.02	0.0043	69.9
10.3	0.02	0.006	70.9	15.2	0.02	0.006	70.5	20.2	0.02	0.006	70.2	25.1	0.02	0.006	69.8	30.2	0.02	0.006	69.2
10.3	0.03	0	76.5	15.3	0.02	0.0075	69.9	20.2	0.02	0.0075	69.5	25.1	0.02	0.0075	68.9	30.1	0.02	0.0075	68.6
10.3	0.04	0	77.2	15.2	0.03	0	75.9	20.2	0.03	0	75.3	25.2	0.03	0	74.8	30.1	0.03	0	74.1
10.2	0.04	0.0015	75.3	15.3	0.04	0	76.7	20.2	0.04	0	76.1	25.2	0.04	0	75.5	30.2	0.04	0	75
10.3	0.04	0.003	73.5	15.2	0.04	0.0015	74.9	20.2	0.04	0.0015	74.4	25.2	0.04	0.0015	73.9	30.1	0.04	0.0015	73.4
10.3	0.04	0.0045	72.7	15.2	0.04	0.003	73	20.2	0.04	0.003	72.5	25.2	0.04	0.003	72.1	30.1	0.04	0.003	71.6
10.3	0.06	0	79.3	15.2	0.04	0.0045	72.4	20.2	0.04	0.0045	71.9	25.1	0.04	0.0045	71.5	30.2	0.04	0.0045	71.1
10.3	0.06	0.0015	76.6	15.2	0.06	0	78.8	20.2	0.04	0.006	71.2	25.2	0.04	0.006	70.7	30.2	0.04	0.006	70.3
10.3	0.06	0.003	74.7	15.2	0.06	0.0015	76.2	20.2	0.06	0	78.3	25.2	0.04	0.0075	70	30.2	0.04	0.0075	69.6
10.2	0.06	0.0045	73.5	15.2	0.06	0.003	74.2	20.2	0.06	0.0015	75.7	25.2	0.06	0	77.7	30.1	0.06	0	77.2
				15.2	0.06	0.0045	73.1	20.2	0.06	0.003	73.8	25.2	0.06	0.0015	75.2	30.2	0.06	0.0015	74.8
								20.2	0.06	0.0045	72.7	25.1	0.06	0.003	73.5	30.1	0.06	0.003	73.1
								20.2	0.06	0.006	71.7	25.1	0.06	0.0045	72.4	30.1	0.06	0.0045	72
								20.2	0.06	0.006	71.7	25.2	0.06	0.006	71.5	30.2	0.06	0.006	71.1
												25.2	0.06	0.006	71.4	30.1	0.06	0.006	71.1
												25.2	0.06	0.0075	70.7	30.2	0.06	0.0075	70.2

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**Table 3.** Parameters for the Eq. (4).

Parameter	Parameter
$U_{12}-U_{11}$	$-4345.7 \left( \frac{\partial(U_{12}-U_{11})}{\partial A} \right)_{T,P,X}$
$U_{13}-U_{11}$	$-10\,803 \left( \frac{\partial(U_{13}-U_{11})}{\partial A} \right)_{T,P,X}$
$U_{21}-U_{22}$	$4175 \left( \frac{\partial(U_{21}-U_{22})}{\partial A} \right)_{T,P,X}$
$U_{23}-U_{22}$	$-8822.1 \left( \frac{\partial(U_{23}-U_{22})}{\partial A} \right)_{T,P,X}$
$U_{31}-U_{33}$	$-467.93 \left( \frac{\partial(U_{31}-U_{33})}{\partial A} \right)_{T,P,X}$
$U_{32}-U_{33}$	$22\,980 \left( \frac{\partial(U_{32}-U_{33})}{\partial A} \right)_{T,P,X}$

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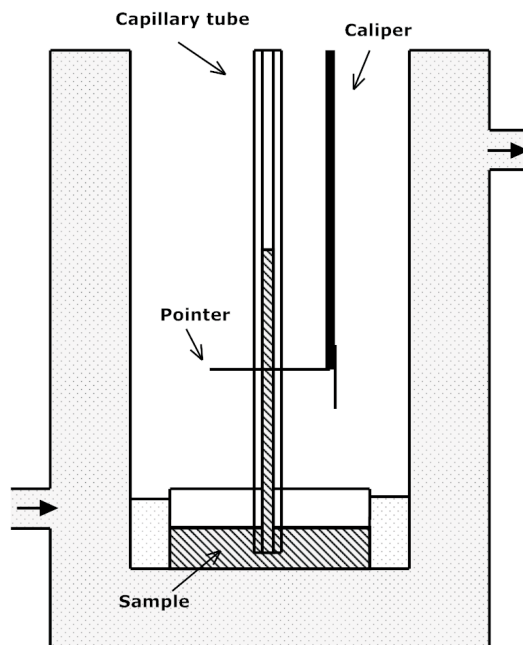
**Table 4.** Modes of the particle size distribution used in the cloud model simulations.

Mode	Geometric mean diameter (nm)	Geometric standard deviation	Number of particles ( $\text{cm}^{-3}$ )
Aitken	45	1.5	300
Accumulation	170	1.5	200

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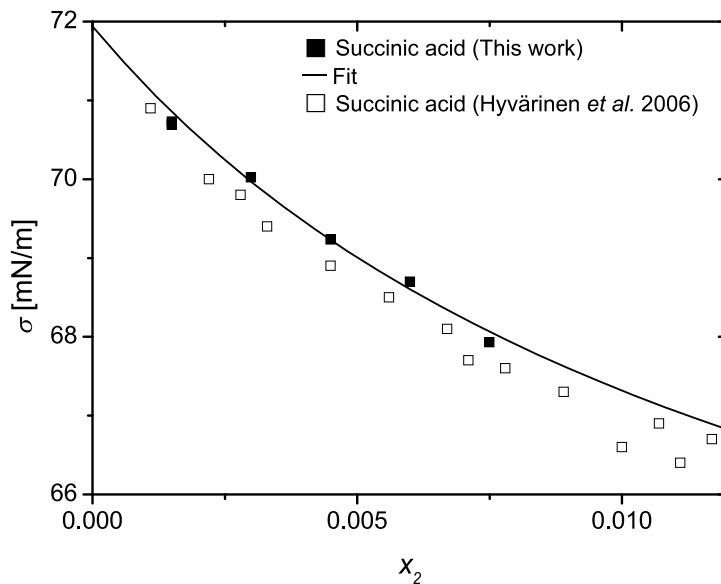
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**Fig. 1.** Experimental setup.[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Printer-friendly Version](#)[Interactive Discussion](#)

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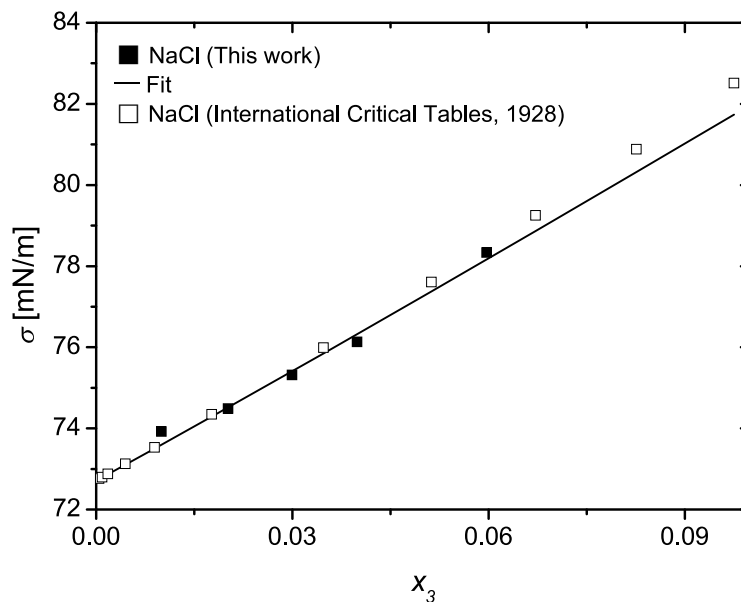


**Fig. 2.** Surface tensions of aqueous succinic acid as a function of the mole fraction of succinic acid ( $x_2$ ) at 25°C.

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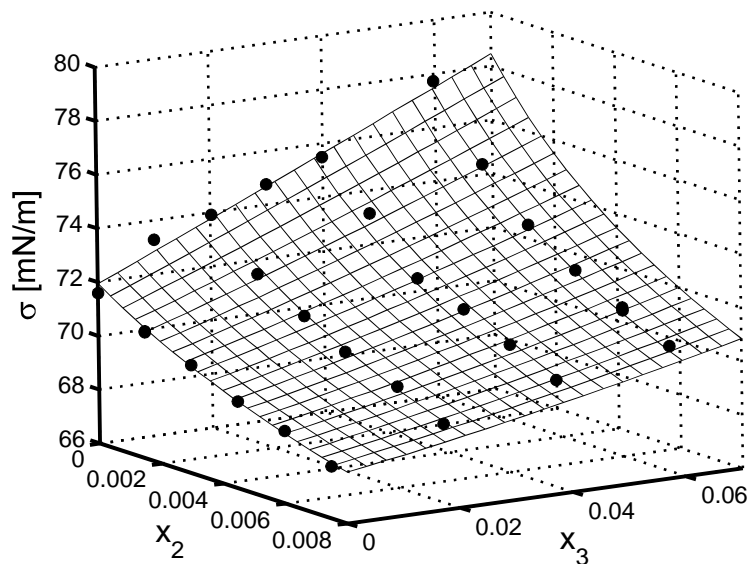


**Fig. 3.** Surface tension of aqueous sodium chloride as a function of the mole fraction of sodium chloride ( $x_3$ ).

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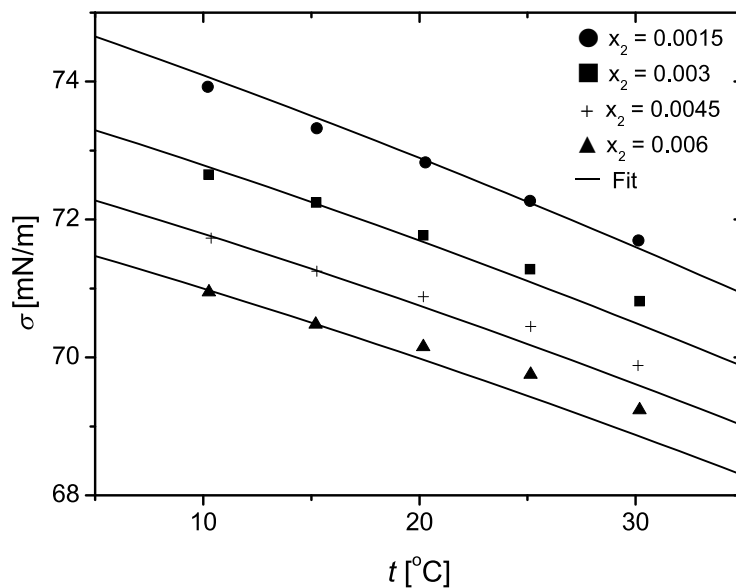
**Fig. 4.** Surface tension of the ternary solution of sodium chloride, succinic acid and water as a function of mole fractions of sodium chloride ( $x_3$ ) and succinic acid ( $x_2$ ) at 25°C. Black dots are the measured surface tensions and grid is the fit to the data by Eq. (4).

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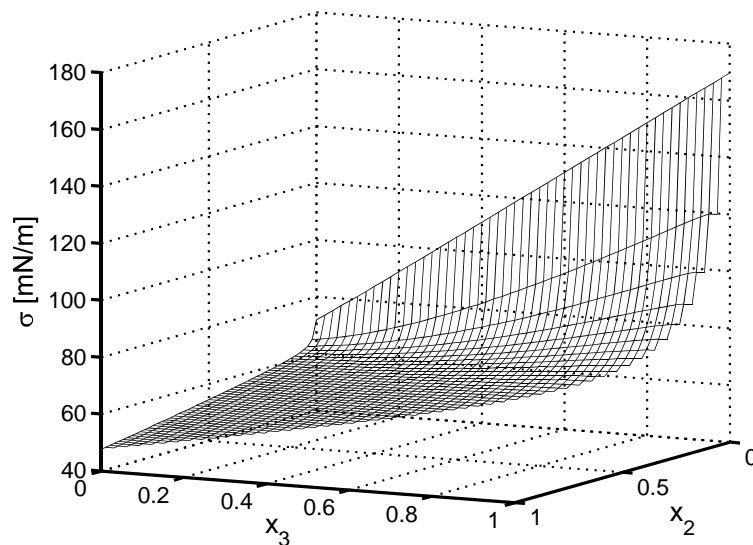


**Fig. 5.** Surface tension of ternary solution of sodium chloride, succinic acid and water as a function of temperature. Mole fraction of sodium chloride is kept constant ( $x_3=0.02$ ).

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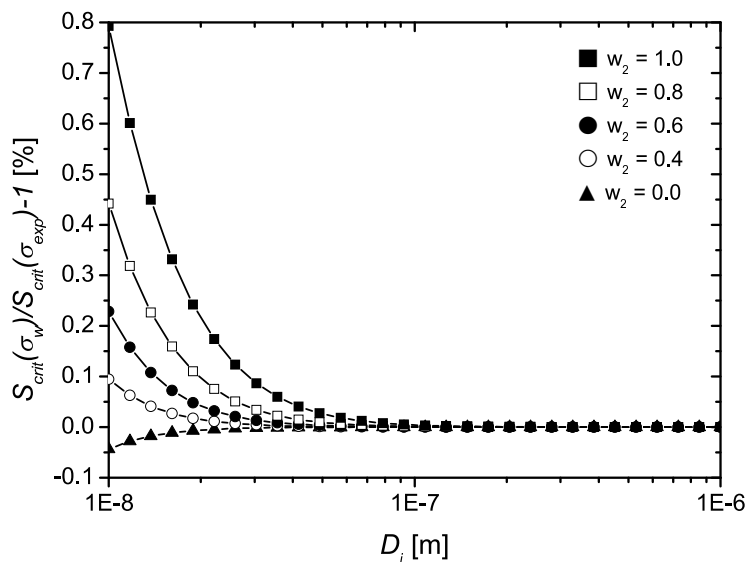


**Fig. 6.** The behaviour of the surface tension parameterization (Eq. 4) as a function of mole fractions ( $x$ ) of sodium chloride and succinic acid at 25°C.

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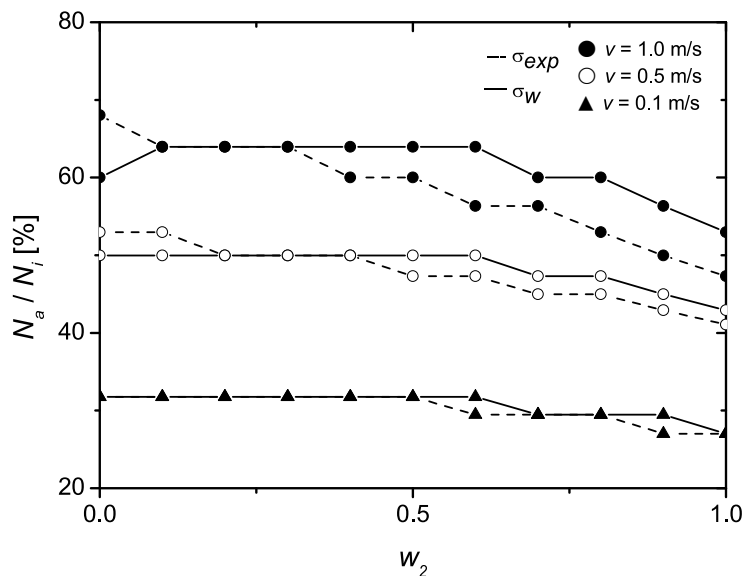


**Fig. 7.** Relative difference in critical supersaturations ( $S_{crit}$ ) obtained using different surface tensions.  $w_2$  is the mass fraction of succinic acid,  $\sigma_w$  is the surface tension of pure water and  $\sigma_{exp}$  is the surface tension parameterization.

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**Fig. 8.** Activated fraction as a function of mass fraction of succinic acid ( $w_2$ ).  $v$  is the updraft velocity,  $\sigma_w$  is the surface tension of pure water and  $\sigma_{exp}$  is the surface tension parameterization.

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