

# Molecular characterization of aerosol-derived water soluble organic carbon using ultrahigh resolution electrospray ionization Fourier transform ion cyclotron resonance mass spectrometry

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## Abstract

Despite the acknowledged relevance of aerosol water-soluble organic carbon (WSOC) to climate and biogeochemical cycling, characterization of aerosol WSOC has been limited. Electrospray ionization Fourier transform ion cyclotron resonance mass spectrometry (ESI FT-ICR MS) is utilized in the present study to provide detailed molecular-level characterization of the high molecular weight (HMW;  $m/z > 223$ ) component of aerosol-derived WSOC collected from rural sites in Virginia and New York, USA. More than 3000 organic compounds were detected by ESI FT-ICR MS within a  $m/z$  range of 223–600 for each sample. Approximately 86% (Virginia) and 78% (New York) of these peaks were assigned molecular formulas using only carbon (C), hydrogen (H), oxygen (O), nitrogen (N), and sulfur (S) as elemental constituents. H/C and O/C molar ratios were plotted on van Krevelen diagrams and indicated a strong contribution of lignin-like and lipid-like compounds to the aerosol-derived WSOC samples. Double bond equivalents were calculated from the molecular formulas and used to identify black carbon (BC) compounds present in aerosol WSOC. BC compounds were found to comprise only 1–4% of the identified compounds in the aerosol-derived WSOC. Several high magnitude peaks in the mass spectra of both samples corresponded to molecular formulas consistent with molecular formulas proposed in previous secondary organic aerosol (SOA) laboratory investigations indicating that SOAs are important constituents of the WSOC. Overall, ESI FT-ICR MS provides the level of molecular characterization needed for detailed compositional and source information of the high molecular weight constituents of aerosol-derived WSOC.

## 1 Introduction

The importance of atmospheric aerosols to several areas of environmental study has been well-documented. Natural and anthropogenically-derived aerosols alter Earth's radiative heat balance, and therefore climate, through scattering and absorption of so-

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lar radiation (e.g. Ramanathan et al., 2001; Satheesh and Moorthy, 2005; Highwood and Kinnersley, 2006). Elevated concentrations of aerosols (specifically hygroscopic aerosols) due to human activities increase the number of cloud condensation nuclei (CCN) that act as seed for cloud droplets. Because of the limited amount of atmospheric water vapor available for cloud formation, an increase in CCN number may reduce the average size of CCN such that it may limit precipitation and thereby increase the lifetime of clouds, thus serving as an indirect positive feedback on climate change (Toon, 2000; Ramanathan et al., 2001; Lohmann and Feichter, 2005).

In addition to the general role of aerosols in climate, fossil fuel and biomass combustion produce anthropogenically-derived aerosols that are known to impair visibility (Charlson, 1969; Jacobson et al., 2000), contribute to ecosystem-level problems via rain acidification (Likens and Bormann, 1974; Driscoll et al., 2001 and references therein) and the transport and deposition of persistent organic pollutants (Dickhut et al., 2000; Galiulin et al., 2002; Jurado et al., 2004), and cause cardiovascular and respiratory problems in humans (Davidson et al., 2005; Highwood and Kinnersley, 2006). Furthermore, atmospherically-derived materials in aerosol form are potentially important in a biogeochemical context. For example, recent studies estimate that between 30 and 90 Tgyr<sup>-1</sup> of aerosol-derived organic carbon (OC; Koch, 2001; Bond et al., 2004) and 8 and 24 Tgyr<sup>-1</sup> black carbon (BC; Penner et al., 1993; Bond et al., 2004) are deposited globally. These fluxes are potentially significant in the context of carbon cycling and budgets at the atmosphere-land-water interfaces, especially in areas where industrial sources are significant. Given the potential quantitative importance of aerosol OC to different terrestrial and aquatic systems, molecular-level characterization of aerosols is critical for both tracing the sources of aerosol OC and assessing its transformations before and after deposition.

Aerosols tend to be highly carbonaceous in nature with OC often comprising 10–30% of total aerosol mass (e.g. Wolff et al., 1986; Jacobson et al., 2000; Tanner et al., 2004; Liu et al., 2005). In addition, as much as 20–70% of aerosol OC is water-soluble (WSOC; Krivacsy et al., 2001; Kleefeld et al., 2002; Yang et al., 2004; Decesari et al.,

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2007). As noted above, only hygroscopic aerosols can act as CCN, making WSOC an important indirect climate agent (Saxena and Hildemann, 1996; Fuzzi et al., 2001; Satheesh and Moorthy, 2005). Aerosol WSOC is also of great potential interest in the context of OC cycling between atmosphere, land, and natural waters because it is likely to be the fraction of aerosol OC that is most rapidly transported along with surface and ground waters through watersheds to lakes, rivers, and estuaries on timescales relevant to carbon biogeochemical cycling.

Despite the potential importance of aerosol WSOC, detailed molecular characterization of the WSOC component of aerosols has thus far been limited. Attempts to characterize WSOC at the molecular level using gas chromatography-mass spectrometry (GC-MS; Mayol-Bracero et al., 2002; Wang et al., 2006) and a combination of ion chromatography and high performance liquid chromatography (HPLC; Yang et al., 2004) characterized less than 10% and 20% of WSOC, respectively. Characterization of aerosol WSOC at the functional group level using HPLC (Mayol-Bracero et al., 2002),  $^1\text{H}$  (Decesari et al., 2000) and cross-polarization-magic angle spinning  $^{13}\text{C}$  (Duarte et al., 2005; Sannigrahi et al., 2006) nuclear magnetic resonance (NMR) spectroscopy, Fourier transform infrared spectroscopy (Duarte et al., 2005), and size exclusion chromatography (Sullivan and Weber, 2006) generally agree with the limited molecular-level investigations (Mayol-Bracero et al., 2002; Yang et al., 2004; Yu et al., 2005; Wang et al., 2006) in identifying mono- and di-carboxylic acids as well as polyconjugated acids (sometimes described as humic-like substances, HULIS) as the most prevalent compounds in WSOC, followed by neutral compounds such as sugars.

The high concentration of acidic species in aerosol-derived WSOC likely indicates the presence of secondary organic aerosols (SOA) formed from the oxidation of naturally and anthropogenically emitted volatile organic carbon (VOC) precursors (Jaoui et al., 2005; Kanakidou et al., 2005; Sullivan and Weber, 2006). To date, however, much of the work identifying SOA compounds has relied on experimental laboratory investigations (e.g. Forstner et al., 1997; Jang and Kamens, 2001; Kanakidou et al., 2005; Heaton et al., 2007), and very few SOA compounds have been identified in ambient

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aerosol samples (Edney et al., 2003; Tolocka et al., 2004; Jaoui et al., 2005). Comprehensive molecular characterization of WSOC derived from ambient aerosol material will therefore complement studies of SOA formation processes, atmosphere-land-water biogeochemical fluxes, and climate-related effects of WSOC.

5 Electro spray ionization coupled with Fourier transform ion cyclotron resonance mass spectrometry (ESI FT-ICR MS) provides detailed molecular characterization of organic matter due to its extremely high resolution and mass accuracies (Marshall et al., 1998; Kujawinski et al., 2002a; Sleighter and Hatcher, 2007). ESI is a “soft” ionization technique that produces minimal fragmentation of the analytes, thus allowing for detection  
10 of intact molecules (Stenson et al., 2002) and is a particularly effective technique for ionizing polar, hydrophilic molecules (Gaskell, 1997; Kujawinski, 2002; Sleighter and Hatcher, 2007) similar to those found in aerosol WSOC. FT-ICR MS provides ultrahigh mass resolving powers (>300 000) and mass accuracy (<1 ppm), enabling the identification of many distinct compounds at a given nominal mass (Marshall et al., 1998;  
15 Kujawinski et al., 2002a; Sleighter and Hatcher, 2007).

Recently, ESI FT-ICR MS has been used to successfully characterize complex mixtures of organic matter including fulvic and humic acids extracted from rivers (Kujawinski et al., 2002a; Stenson et al., 2003; Kujawinski et al., 2004), soils (Kujawinski et al., 2002a; Kramer et al., 2004; Hockaday et al., 2006), peat (Kramer et al., 2004), and degraded  
20 wood (Kujawinski et al., 2002a; Kim et al., 2003a; Kramer et al., 2004) as well as dissolved organic matter (DOM) from riverine (Kim et al., 2003a,b; Stenson et al., 2003; Kim et al., 2004) and marine waters (Koch et al., 2005; Tremblay et al., 2007), ice cores (Grannas et al., 2006), mangrove porewaters (Koch et al., 2005; Tremblay et al., 2007), and soils (Hockaday et al., 2006). Amongst the key findings in recent  
25 ESI FT-ICR MS studies has been the detection of potential BC compounds (Kim et al., 2003b; Kim et al., 2004; Kramer et al., 2004; Hockaday et al., 2006). With biomass and fossil fuel combustion thought to be the dominant sources of BC, ESI FT-ICR MS analysis of aerosol WSOC may be useful for detecting the mobile, water-soluble component of BC present in aerosols. The recent studies of Grannas et al. (2006) have

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shown that BC compounds can be identified in melt water from ice cores and represent historically-deposited atmospheric precipitation in snow-accumulating regions of the world. In the present study, we demonstrate the utility of ESI FT-ICR MS for characterizing aerosol-derived WSOC and present novel information on the molecular composition of this material collected in two different watersheds of the northeastern United States.

## 2 Experimental section

### 2.1 Sample collection and field methods

Two-day integrated high-volume aerosol samples ( $>4000\text{ m}^3$ ) were collected during 16–18 August 2006 at the Institute of Ecosystem Studies Environmental Monitoring Station in Millbrook, NY ([http://www.ecostudies.org/emp\\_purp.html](http://www.ecostudies.org/emp_purp.html)) and 7–9 November 2006 at the National Atmospheric Deposition Program (NADP) site (VA98) located in Gloucester County, VA (<http://nadp.sws.uiuc.edu/sites/siteinfo.asp?net=NTN\&id=VA98>) using high-volume total suspended particulate (TSP) air samplers (Model GS2310, ThermoAndersen, Smyrna, GA). Both sites are located in rural environments and are more than 30 km from major industrial emissions. Air was drawn through pre-ashed (3 h,  $525^\circ\text{C}$ ) and pre-weighed high-purity quartz microfibre filters (20.3 cm  $\times$  25.4 cm, nominal pore size  $0.6\ \mu\text{m}$ ; Whatman QM-A grade) for collection of aerosol particles. Following collection, aerosol filter samples were transferred to pre-ashed (3 h,  $525^\circ\text{C}$ ) aluminum foil pouches and stored in the dark in a carefully cleaned air-tight polycarbonate desiccator until analysis.

### 2.2 Aerosol-derived WSOC $\text{C}_{18}$ extraction procedure

Approximately half of each aerosol filter was cut into strips using solvent-cleaned (hexane, acetone, and methanol) razor blades and placed in pre-combusted ( $500^\circ\text{C}$ )

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and solvent-cleaned 1-L glass beakers. Approximately 200 mL of LC-MS grade water (Fisher Scientific) was added to the filter strips, and samples were sonicated for 30 min to extract the WSOC from the filters. The extracted organic matter was then isolated and concentrated from the WSOC filtrates using C<sub>18</sub> solid phase extraction disks (3M, Empore) following previously established protocols (Kim et al., 2003a). The C<sub>18</sub> disks were activated using LC-MS-grade water and methanol (Fisher Scientific), and each WSOC sample was acidified to a pH of 2 with 10 M HCl before passing through the disk. The sorbed material was rinsed with LC-MS grade water before eluting it off the disk with 4–6 mL of LC-MS grade methanol. Due to the qualitative nature of these studies, the recovery from the C<sub>18</sub> disk was not measured; however, previous studies have shown that approximately 42–60% of freshwater dissolved organic matter is recovered by this technique (Louchouart et al., 2000; Kim et al., 2003a).

### 3 Analytical methods

Because previous studies have determined that water/methanol mixtures yield higher quality mass spectra (Kujawinski et al., 2002b; Rostad and Leenheer, 2004), the C<sub>18</sub>-extracts were diluted by 25% with LC-MS grade water. In order to increase the ionization efficiency, ammonium hydroxide was added immediately prior to ESI, raising the pH of the sample to approximately 8. Samples were continuously infused into the Apollo II ESI ion source of a Bruker Daltonics 12 Tesla Apex Qe FT-ICR MS, housed at the College of Sciences Major Instrumentation Cluster (COSMIC) at Old Dominion University (<http://www.sci.odu.edu/sci/cosmic/index.shtml>). Samples were introduced by a syringe pump providing an infusion rate of 120  $\mu$ L/h. All samples were analyzed in negative ion mode, and electrospray voltages were optimized for each sample. Previous studies have shown that the negative ion mode avoids the complications associated with the positive ion mode in which alkali metal adducts, mainly Na<sup>+</sup>, are observed along with protonated ions (Brown and Rice, 2000; Rostad and Leenheer, 2004). Ions were accumulated in a quadrupole ion trap for 1.0 s before being transferred to the ICR

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cell. Exactly 300 transients, collected with a 4 MWord time domain, were added, giving about a 30 min total run time for each sample. The summed free induction decay (FID) signal was zero-filled once and Sine-Bell apodized prior to fast Fourier transformation and magnitude calculation using the Bruker Daltonics Data Analysis software.

### 5 3.1 Mass calibration and molecular formula assignments of aerosol WSOC

Prior to data analysis, all samples were externally calibrated with an arginine cluster standard and internally calibrated with fatty acids naturally present within the sample (Sleighter et al., 2008). The ultrahigh resolving power of 12 T FT-ICR MS is capable of separating  $m/z$  values to a mass accuracy of less than 1 ppm. A molecular formula calculator developed at the National High Magnetic Field Laboratory in Tallahassee, FL (Molecular Formula Calc v.1.0 ©NHMFL, 1998; <http://www.magnet.fsu.edu/>) generated empirical formula matches using carbon (C), hydrogen (H), oxygen (O), nitrogen (N), sulfur (S), and phosphorus (P). Only  $m/z$  values with a signal-to-noise above 4 were inserted into the molecular formula calculator. In the vast majority of cases, the exact mass of each assigned formula agreed with the  $m/z$  value to within less than 0.5 ppm.

#### 3.1.1 Data processing

Molecular formulas generated by the molecular formula calculator were pre-processed using a MatLab file (The MathWorks Inc., Natick, MA) that employed several conservative rules in order to eliminate compounds not likely to be observed in nature. The pre-processing file eliminated all molecular formulas in which:  $O/C \geq 1.2$ ,  $H/C \geq 2.25$ ,  $H/C \leq 0.3$ ,  $N/C \geq 0.5$ ,  $S/C \geq 0.2$ ,  $P/C \geq 0.1$ ,  $(S+P)/C \geq 0.2$ , and DBE (double bond equivalents)  $< 0$ , where DBE is calculated as follows:

$$DBE = (2c + 2 - h + n + p)/2 \quad (1)$$

for any formula  $C_cH_hN_nO_oS_sP_p$ . DBE indicates the number of rings and double bonds in a molecule and is a measure of the degree of unsaturation in a given compound (Hockaday et al., 2006).

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In addition to these rules, the MatLab file searched for and eliminated peaks detected by the FT-ICR MS that were 1.003 mass units greater than another detected peak, under the assumption that the two peaks represent the same compound, with the peak at the higher  $m/z$  having a  $^{13}\text{C}$  in place of a  $^{12}\text{C}$ . Following pre-processing, molecular formulas were assigned following the “formula extension” approach described by Kujawinski and Behn (2006). Phosphorous is typically not a quantitatively significant component of atmospheric materials (Chen et al., 2002; Grimshaw and Dolske, 2002; Baker et al., 2006); therefore all molecular formulas containing phosphorous were eliminated for ease of processing. Additionally, by analyzing solvent blanks, peaks below  $m/z$  223 were determined to be high-frequency noise and were not analyzed.

## 4 Results and discussion

### 4.1 Mass spectra

ESI FT-ICR mass spectra for aerosol-derived WSOC samples from both locations showed over 3000 peaks and remarkably similar patterns in peak distribution. The greatest peak magnitudes were located between  $m/z$  250 and 375, and considerably smaller magnitude peak clusters were centered around  $m/z$  450 and 550 (Fig. 1a and b). Peak magnitude is not necessarily indicative of a compound’s concentration in a sample due to inherent biases of  $\text{C}_{18}$  extractions and electrospray ionization efficiencies (Hockaday et al., 2006). In addition, only two samples were investigated here, so further work is needed to determine whether this is a characteristic molecular weight distribution of aerosol WSOC. Nonetheless, the similarity between the spectra is striking given the differences in dates and locations of sampling (August 2006, Millbrook, NY vs. November 2006, Harcum, VA). Dismal Swamp DOM (Sleighter and Hatcher, 2007) and Mount Rainier humic acid (Kujawinski et al., 2002a) also showed strongest peak magnitudes at  $m/z$  250 to 375 in their ESI FT-ICR spectra, indicating similar molecular weight distributions between these three different sample types.

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Spectral details for both samples (see insets for Fig. 1a and b) show that  $m/z$  330–340 demonstrated a typical pattern of several peaks at low mass defects (mass defect is the distance a peak is displaced from the exact nominal mass) from a given odd-numbered nominal mass, and fewer, less intense peaks at even-numbered nominal masses (e.g. Kujawinski, 2002; Kim et al., 2004; Sleighter and Hatcher, 2007). Even-numbered nominal mass peaks are indicative of either  $^{13}\text{C}$  isotopic peaks (observed at a mass difference of  $1.003 m/z$  greater than its  $^{12}\text{C}$  counterpart) or N-containing compounds with an odd number of N (Koch et al., 2005). N-containing compounds are preferentially discriminated against in  $\text{C}_{18}$  extraction methods because the polar nature of organic nitrogen precludes its quantitative retention on the hydrophobic  $\text{C}_{18}$  disk (Benner, 2002; Koch et al., 2005); therefore, any N-containing compounds present in the WSOC samples would likely be present at relatively smaller peak magnitudes.

Figure 2 illustrates both the ultrahigh mass resolving power of FT-ICR MS and the heterogeneous nature of aerosol WSOC. Similar to other mixtures of natural organic matter (e.g. Kim et al., 2004; Kujawinski et al., 2004; Koch et al., 2005), aerosol WSOC in this study was found to be highly complex, containing more than 3000 compounds in a single sample (Fig. 1) and 10–20 compounds at any given nominal mass (Fig. 2). Less than 0.0001 Da often separated two distinct peaks (i.e. compounds; Fig. 2). Only instrumentation with mass-resolving powers as high as FT-ICR MS can obtain the separation necessary to resolve compounds at such small mass differences.

## 4.2 Formula assignments

Using a formula extension approach similar to that described by Kujawinski and Behn (2006), 86% and 78% of the peaks identified in the mass spectra for the Virginia and New York aerosol WSOC samples, respectively, were assigned molecular formulas. Most of the identified compounds contained only C, H, and O (Table 1). The New York aerosol WSOC had a higher percentage of compounds containing C, H, O, and S (26% versus 17%; Table 1) than the Virginia sample. In contrast, compounds containing C, H, O, and N accounted for 26% of those identified in the Virginia sample vs. only 16%

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in the New York sample (Table 1). While ESI FT-ICR mass spectra do not provide a quantitative estimate of compound concentrations, the New York and Virginia WSOC samples were analyzed on the same instrument using the same protocols and methods, and therefore the resulting spectra reflect the same inherent biases. Because of the similar sample processing, we may calculate and compare magnitude-weighted percent contributions for various compound groups (e.g. C-H-O, C-H-O-N, C-H-O-S, C-H-O-N-S; Table 1). These magnitude-weighted percent contributions further support a greater contribution of S-containing compounds to New York aerosol WSOC than to Virginia WSOC, and a greater contribution of N-containing compounds to Virginia aerosol WSOC compared to that from New York.

### 4.3 Van Krevelen analysis

The identified molecular formulas of aerosol WSOC are represented in van Krevelen diagrams (Fig. 3a and b) which plot molar ratios (H/C vs. O/C) of compounds against each other. Also plotted are representative molar ratios for major known classes of natural and anthropogenic organic compounds (i.e. ovals in Fig. 3a and b) in order to provide potential source information for the aerosol WSOC samples (see also Kim et al., 2003b; Kujawinski et al., 2004; Sleighter and Hatcher, 2007). In addition to major compound classes frequently used to apportion sources of dissolved organic matter (Kim et al., 2003b; Kujawinski et al., 2004; Sleighter and Hatcher, 2007), SOAs and mono- and di-carboxylic acids, compounds frequently associated with WSOC, are included to provide additional likely sources to aerosols (Fig. 3a and b; Yang et al., 2004; Yu et al., 2005; Sullivan and Weber, 2006; Wang et al., 2006).

The majority of aerosol WSOC compounds identified by ESI FT-ICR MS in this study had molar H/C and O/C ratios similar to lipids, lignin, mono- and dicarboxylic acids, and SOAs (Fig. 3a and b). Compounds with molar ratios reflective of protein sources were also abundant, however, the majority of these compounds do not contain nitrogen, a component of every amino acid, (Fig. 3a and b) indicating that proteins are an unlikely source for these particular compounds in the present study. Few compounds

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plotted within the regions characteristic of tannins, cellulose, and condensed hydrocarbons (Fig. 3a and b) suggesting that these groups are also not major contributors to aerosol WSOC. Many of the compounds plotted outside any of the source signature regions and had low O/C ( $\sim 0$  to 0.2) and H/C ( $\sim 1$  to 1.5) indicating an unknown source.

5 There are currently very few reports of SOA H/C and O/C molar ratios (Reinhardt et al., 2007; Altieri et al., 2008), however, with further investigation the region representing SOA may prove to be more extensive and encompass those compounds not presently represented by any known sources.

10 The relatively broad distributions of H/C and O/C molar ratios for compounds in these samples (Fig. 3a and b) again illustrate the complex nature of aerosol WSOC, however, certain patterns emerge. The majority of S-containing compounds identified in both samples had H/C values greater than 1.2 and simultaneously showed high O/C ratios. Organosulfur compounds have been detected in previous aerosol WSOC studies (Romero and Oehme, 2005; Gilardoni et al., 2007). Romero and Oehme (2005) 15 further observed  $\text{HSO}_4^-$  mass fragments in high molecular weight (HMW) atmospheric HULIS and postulated that these anions were covalently bound by sulfonation or sulfation processes. The high O/C ratios of many S-containing compounds observed in the New York and Virginia aerosol WSOC samples (Fig. 3a and b) are also consistent with covalently bound  $\text{HSO}_4^-$ . Elevated H/C molar ratios ( $>1.5$ ) are indicative of saturated hydrocarbons with few double bonds, and the high H/C ratios in S-containing compounds in these samples indicate that any sulfonation or sulfation processes resulted in mostly saturated compounds. S-containing aromatic compounds that would show much lower H/C ratios are not evident in these samples. In contrast, N-containing and C-H-O compounds (Fig. 3a and b) frequently have H/C values  $<1$  and do not typically 20 have O/C values  $>0.6$ , suggesting that the nitrogenous WSOC compounds in these samples tended to be highly carbonaceous, condensed compounds. Previous laboratory studies of SOAs have also reported the formation of nitro-aromatic compounds from the photooxidation of aromatic compounds in the presence of  $\text{NO}_x$  (Forstner et al., 1997; Jang and Kamens, 2001; Alfarrá et al., 2006). The data presented here are 25

consistent with the presence of nitro-aromatic compounds as well.

#### 4.4 Black carbon in aerosol-derived WSOC

The aerosol WSOC samples from New York and Virginia contained several compounds with DBE/C values greater than 0.7, a characteristic of condensed aromatic ring structures and a cut-off value proposed for the identification of BC compounds (Fig. 4; Hockaday et al., 2006). BC compounds defined in this manner made up only 4% and 1% of the identified compounds in Virginia and New York aerosol WSOC, respectively, and were present at small magnitudes relative to the majority of other compounds present (Fig. 4a, b). When peak magnitudes were accounted for as in Table 1 above, BC compounds accounted for only 1.5% (Virginia) and 0.3% (New York) of the total peak magnitudes.

BC has traditionally been studied in particulate OM (e.g. Mitra et al., 2002; Gatari and Bowman, 2003; Dickens et al., 2004). However, BC may become hydrophilic in the course of its oxidation (Kamegawa et al., 2002; Park et al., 2005; Zuberi et al., 2005), and several studies of aqueous OM mixtures have identified a BC component (Mannino and Harvey, 2004; Kim et al., 2004; Kramer et al., 2004; Hockaday et al., 2006). A recent FT-ICR MS study of freshwater DOM identified BC using molar H/C and O/C ratios using a similar approach to the one employed in the present study but did not report the number of peaks characterized as BC (Kim et al., 2004). A study of BC in DOM from the Delaware Bay found that 9% of bay DOC and 4–7% of coastal ocean DOC was BC (Mannino and Harvey, 2004). The authors listed sediment resuspension and atmospheric transport from nearby Philadelphia, PA as likely sources of BC to the bay. We are unaware of aerosol WSOC studies that have quantified BC, but soot oxidation has been demonstrated to form WSOC compounds (Decesari et al., 2002). While the relative paucity of BC compounds identified in this study does not support a strong aerosol WSOC source for BC to riverine and coastal DOC, the presence of BC compounds in WSOC from both of these rural sites suggests that areas having stronger BC sources such as urban regions may contribute greater amounts of BC to

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riverine and coastal DOC.

Radiocarbon analysis of the New York aerosol WSOC from August 2006 showed a mean  $\Delta^{14}\text{C}$  signature of  $-230\text{‰}$  ( $n=3$ ; Wozniak et al., 2008). A simple two-source isotopic mass balance assuming one source devoid of  $^{14}\text{C}$  (e.g. fossil fuels) and another source having present-day levels of  $^{14}\text{C}$  (e.g. modern living biomass) suggests that more than 20% of the New York WSOC comes from a fossil source. BC emitted as a byproduct of fossil fuel combustion represents a logical potential source of this aged WSOC. The data presented here, however, do not support a significant input of BC to aerosol WSOC, and therefore other sources of aged organic matter, both natural and anthropogenic, may be responsible for the aged WSOC (e.g. aged soil organic matter, SOAs from fossil fuel precursors, etc.). In addition, several of the identified BC compounds contain N in their molecular formulas, and as discussed previously,  $\text{C}_{18}$  extraction does not retain organic N compounds efficiently (Benner, 2002; Koch et al., 2005). Therefore, BC-derived compounds present in the initial WSOC sample may not be quantitatively represented as well as non-N containing compounds in the FT-ICR mass spectra. Alternately,  $\text{DBE}/\text{C} \geq 0.7$  may be too conservative as a cut-off for a complete identification of BC compounds (Fig. 4a and b; Hockaday et al., 2006).

While BC compounds may comprise only a small portion of identified molecular formulas in the aerosol WSOC samples analyzed here, their identification nonetheless highlights another application of ESI FT-ICR MS. BC is generally defined as carbonaceous material thought to be composed of a highly refractory, slow-cycling pool of compounds resulting from combustion processes with relevance to climate and carbon cycling issues and can be a significant portion of aerosol carbonaceous material (e.g. Novakov et al., 2005; and references therein). In a biogeochemical context, the identification of BC in aerosol WSOC suggests that BC may become desorbed into rainwater and transported through watersheds to various aquatic systems. To this point, BC has primarily been studied using one of several operational definitions that do not measure the full spectrum of BC (Masiello, 2004; Hammes et al., 2007). The use of ESI FT-ICR MS to identify BC in aerosols may therefore provide molecular level information

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allowing for better characterization of BC in WSOC.

#### 4.5 Potential contributions of secondary organic aerosols to aerosol-derived WSOC

Formula assignments for many of the FT-ICR MS peaks in this study were consistent with formulas proposed in experimental laboratory investigations of SOA formation by other researchers (Table 2). While molecular structure can not be deduced from the data collected in the present study, the molecular formulas are consistent with the presence of at least certain SOA compounds and illustrate how the extremely high mass resolution of FT-ICR MS may be utilized to identify dominant SOA species in field-collected aerosols, aerosol-WSOC, rainwater and other natural aqueous samples.

Of the molecular formulas in Table 2,  $C_{18}H_{28}O_4$  was the most prevalent potential SOA species in the Virginia sample, while  $C_{20}H_{32}O_4$  was the most prevalent potential SOA species in the New York sample. Heaton et al. (2007) observed  $C_{18}H_{28}O_4$  as a product of  $\beta$ -pinene ozonolysis and suggested its formation is via reaction of a monomer end product and a hydroperoxide intermediate of  $\beta$ -pinene ozonolysis. In comparison,  $C_{20}H_{32}O_4$  (Table 2) was a product of  $\alpha$ -pinene ozonolysis, and its presence was attributed to dimerization of pinonaldehyde, a known product of primary ozonolysis, via either aldol condensation or *gem*-diol formation (Tolocka et al., 2004).

The majority of previous experimental and field studies identifying SOA compounds focused on low molecular weight (LMW) species ( $m/z < 200$ ; e.g. Forstner et al., 1997; Jang and Kamens, 2001). However, recent experimental work also argues for the formation of HMW SOA compounds that are oligomers of precursor compounds (Gao et al., 2004; Kalberer et al., 2004; Tolocka et al., 2004; Dommen et al., 2006; Heaton et al., 2007). The abundance of HMW peaks in the mass spectra (Fig. 1a and b) and molecular formula assignments in Table 2 agree with these recent studies on the contribution of HMW compounds to SOA (Gao et al., 2004; Kalberer et al., 2004; Tolocka et al., 2004; Dommen et al., 2006; Heaton et al., 2007). While the ESI FT-ICR MS method utilized in this study does not characterize LMW SOA products, detailed molecular characterization of HMW SOA using FT-ICR MS is possible. This approach may help

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establish which of the many HMW SOA compounds identified in experimental laboratory investigations are prevalent in field samples, thus helping to identify the dominant formation processes and pathways under natural environmental conditions. Laboratory studies have also employed ESI FT-ICR MS to examine the behavior of  $\alpha$ -pinene in the presence of ozone to reveal that polymerization processes were important in the formation of HMW SOA (Tolocka et al., 2004; Reinhardt et al., 2007). Tolocka et al. (2004) also found several species whose presence could not be explained by monomer polymerization or degradation and proposed unknown complex chemical pathways for their formation. This study also reported SOA oligomers from a field sample illustrating the utility of FT-ICR MS in validating laboratory investigations (Tolocka et al., 2004). ESI FT-ICR MS may alternately be used in field and laboratory investigations to detect SOA compounds that may be abundant but previously unidentified.

## 5 Conclusions

ESI FT-ICR MS is well-adapted to molecularly characterizing aerosol WSOC as a result of its extremely high mass resolving power and ability to ionize and detect polar, hydrophilic molecules such as the HULIS-type molecules in WSOC. Van Krevelen diagrams of compounds found in aerosol WSOC may be further useful for understanding the general compositional features of WSOC compounds. Finally, ESI FT-ICR MS may be particularly well-suited to identifying BC and SOA molecular formulas, giving investigators an inventory of BC and SOA compounds in aerosol WSOC that may permit more detailed study of these highly complex substances for better understanding their formation and degradation pathways. Further work is required to make ESI FT-ICR MS analyses more quantitative if at all possible, but presently has been shown to be an excellent qualitative tool available to the aerosol and atmospheric chemistry communities.

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**Table 1.** Percent occurrence of formula groups and magnitude-weighted percent contribution of formula groups to all peaks assigned molecular formulas in the mass spectra. Formulas are grouped based on their elemental constituents.

Elemental Constituents	Percent Occurrence		Magnitude-Weighted Percent Contribution	
	VA	NY	VA	NY
C, H, O	57	58	77	75
C, H, O, N	26	16	12	6
C, H, O, S	17	26	11	19
C, H, O, N, S	<1	<1	<1	<1

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**Table 2.** Elemental formulas of compounds consistent with proposed or identified SOA compounds from selected previous studies. Relative peak magnitudes are ranked from 1 (strongest peak assigned a molecular formula in spectrum) to  $n$  (weakest peak assigned a molecular formula in spectrum; for VA,  $n=2098$ ; for NY,  $n=2261$ ).

Elemental Formula	$m/z$ of Ion	Relative Peak Magnitude		Proposed Formation Mechanism
		VA	NY	
$C_{11}H_{18}O^a$	245.1031	1924	not present	Dimer product of 1-methyl cyclohexene ozonolysis
$C_{17}H_{26}O^b$	309.1708	40	53	Dimer product of $\beta$ -pinene ozonolysis
$C_{17}H_{26}O^b$	325.1657	45	106	Dimer product of $\beta$ -pinene ozonolysis
$C_{18}H_{28}O^c$	307.1915	11	28	Dimer product of $\beta$ -pinene ozonolysis
$C_{18}H_{28}O^c$	339.1813	42	49	Dimer product of $\beta$ -pinene ozonolysis
$C_{18}H_{28}O^c$	355.1762	115	203	Dimer product of $\beta$ -pinene ozonolysis
$C_{18}H_{30}O^a$	357.1919	155	307	Dimer product of $\alpha$ -pinene ozonolysis
$C_{19}H_{30}O^b$	369.1919	119	143	Dimer product of $\beta$ -pinene ozonolysis
$C_{19}H_{32}O^b$	355.2126	95	103	Dimer product of $\beta$ -pinene ozonolysis
$C_{20}H_{32}O^c$	335.2228	36	7	Dimer product of $\alpha$ -pinene ozonolysis
$C_{20}H_{32}O^c$	351.2177	68	19	Dimer product of $\alpha$ -pinene ozonolysis
$C_{19}H_{30}O^c$	337.2021	32	16	Dimer product of $\alpha$ -pinene ozonolysis
$C_{18}H_{28}O^c$	323.1864	26	22	Dimer product of $\alpha$ -pinene ozonolysis

<sup>a</sup> Gao et al., 2004;

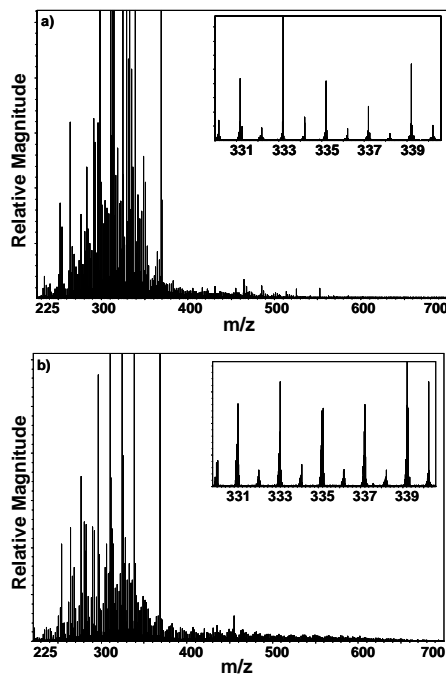
<sup>b</sup> Heaton et al., 2007;

<sup>c</sup> Tolocka et al., 2004.



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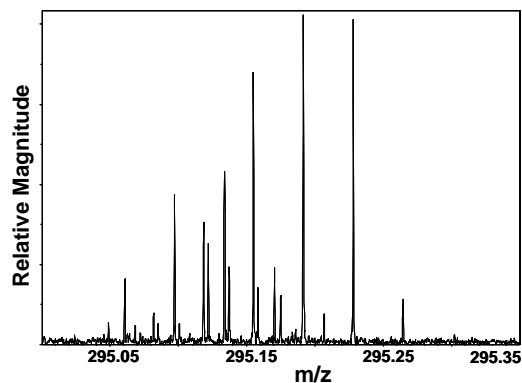


**Fig. 1.** ESI FT-ICR mass spectra for aerosol WSOC samples collected in **(a)** August 2006 in New York and **(b)** November 2006 in Virginia. Insets highlight the  $m/z$  330–340 region showing characteristic peak distributions.

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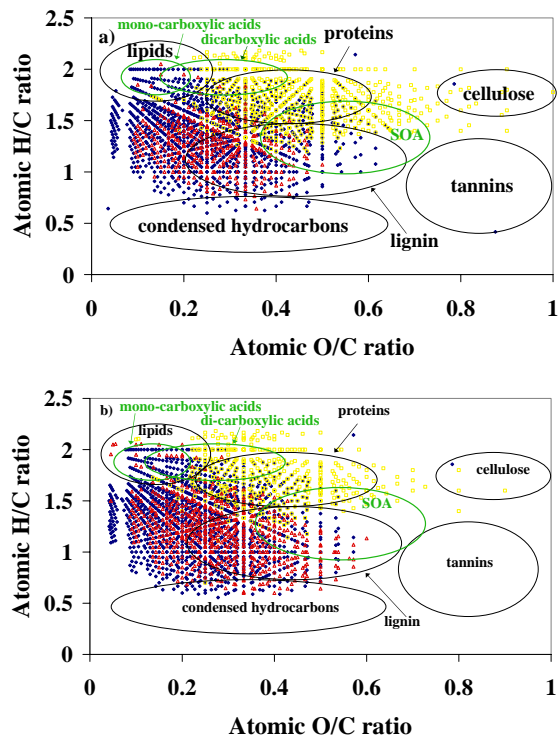


**Fig. 2.** Expanded ESI FT-ICR mass spectrum region at  $m/z$  295 for aerosol WSOC sample from New York in August of 2006.

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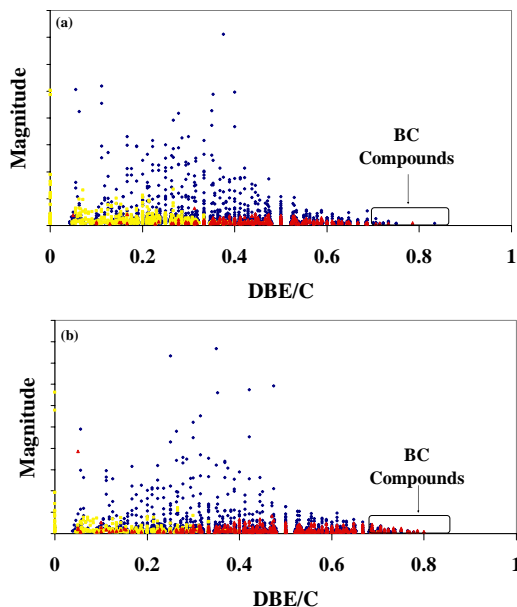


**Fig. 3.** Van Krevelen plots for elemental formulas assigned to FT-ICR mass spectra peaks in aerosol WSOC samples from (a) New York and (b) Virginia. Blue diamonds represent compounds containing only C, H, and O, yellow squares represent S-containing compounds, and red triangles are N-containing compounds. Black ovals represent traditional potential source molecular classes (after Sleighter et al., 2007). The green SOA oval represents published molar H/C and O/C ratios from laboratory investigations of secondary organic aerosols (Reinhardt et al., 2007; Altieri et al., 2008). O/C ratios for the green ovals representing mono- and dicarboxylic acids were calculated for molecules containing 10–30 carbons, and the H/C ratios for these ovals assumed little branching and few double bonds.

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**Fig. 4.** Peak magnitude vs. DBE/C for elemental formulas identified in aerosol WSOC samples from **(a)** New York and **(b)** Virginia. Blue diamonds represent compounds containing only C, H, and O, yellow squares represent S-containing compounds, and red triangles are N-containing compounds. For both samples, the y-axis was augmented to make lower magnitude compounds visible. Several high magnitude C, H, O compounds at DBE/C between 0.3 and 0.5 are not shown.

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